

Stimuli Responsive Ionogels for Sensing Applications—An Overview

Andrew Kavanagh, Robert Byrne, Dermot Diamond and Kevin J. Fraser *

CLARITY; The Centre for Sensor Web Technologies, National Centre for Sensor Research, School of Chemical Sciences, Glasnevin, Dublin 9, Ireland; E-Mails: andrew.kavanagh@dcu.ie (A.K.); Robert.Byrne@dcu.ie (R.B.); dermot.diamond@dcu.ie (D.D.);

* Author to whom correspondence should be addressed; E-Mail: kevin.fraser@dcu.ie; Tel.: +353-1-700-6009; Fax: +353-1-700-5404.

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Abstract: This overview aims to summarise the exiting potential of “*Ionogels*” as a platform to develop stimuli responsive materials. Ionogels are a class of materials that contain an Ionic Liquid (IL) confined within a polymer matrix. Recently defined as “a solid interconnected network spreading throughout a liquid phase”, the ionogel therefore combines the properties of both its solid and liquid components. ILs are low melting salts that exist as liquids composed entirely of cations and anions at or around 100 °C. Important physical properties of these liquids such as viscosity, density, melting point and conductivity can be altered to suit a purpose by choice of the cation / anion. Here we provide an overview to highlight the literature thus far detailing the encapsulation of IL and responsive materials within these polymeric structures. Exciting applications in the areas of optical and electrochemical sensing, solid state electrolytes and actuating materials shall be discussed.

Keywords: Stimuli Responsive Polymers; Ionogels; Ionic Liquids; Hybrid Materials; Molecular Photoswitches; Solid State Electrolytes

1. Introduction

The concept of a chemical sensor is one in which a material is used as a sensing agent and exhibits a selective interaction with a target species or analyte[1,2]. The specific interaction between the sensor and analyte produces a signal, which can then be observed via an appropriate detection scheme [3,4]. Liquid based sensors can suffer from volatility and handling issues, meaning their performance suffers over time. Producing solid-state platforms is of great importance for some applications, as the solid-

state removes many of the issues associated with that of the liquid state[5]. There is great interest therefore in solid-state chemical sensors that can provide reliable signals at a low unit cost, and through careful optimisation of the sensitive polymer composition, prevent leaching or removal of key components over time[6]. Polymer gels have been employed in sensing templates for this purpose and we will explore their use in detail in this review [7,8].

A polymer gel is defined as an interconnected polymer network formed within a liquid phase[9,10]. When the polymer network is generated in the presence of an Ionic Liquid (IL), the resultant gel has been termed an *ionogel* within the literature[11]. Ionogels are therefore a new class of hybrid material that combine the physical properties of both the polymer gel and the physically entrapped IL within[12].

A recent review of the area focused on the differing subclasses of polymers employed (i.e. organic, inorganic and hybrid repeating units), and the modes of preparation of ionogels[13]. Previous reviews focused on the interaction and mobility of the IL within the polymer network[14]. This overview will complement these reviews by focusing on the application of ionogels as functional materials for direct application as sensing and actuation agents. Publications detailing the response of ionogels to changes in pH, metal ion chelation, incident electromagnetic radiation and interactions with biomolecules will be discussed. Our group in particular have detailed the use of ionogels to produce opto/electronic sensors for transition metal ions[15,16], as the basis of electrochromic devices[17], as actuating materials for controlling fluid movement[18] as reference electrodes[19], and building biosensing platforms [20] This overview will present the current status of the use of ionogels in these areas, beginning first with an introduction to the area of ILs themselves.

2. Ionic Liquids

According to current convention, a salt melting below the normal boiling point of water is known as an IL, thus forming liquids that are comprised entirely of cations and anions at room temperature[21]. In contrast to conventional organic liquids / solvents, important physical properties of ILs such as viscosity, density, melting point and conductivity can be tuned to suit a particular need by the appropriate choice of the cation / anion combination[22]. ILs typically contain a large bulky asymmetric cation together with a smaller π -delocalised anion which overwhelmingly exhibit electrostatic interactions; thereby preventing the formation of a structured lattice[23,24]. ILs exhibit good thermal stability[25], are intrinsically conductive[26] and have been shown to have electrochemical windows as high as 5.0 V in some cases[23].

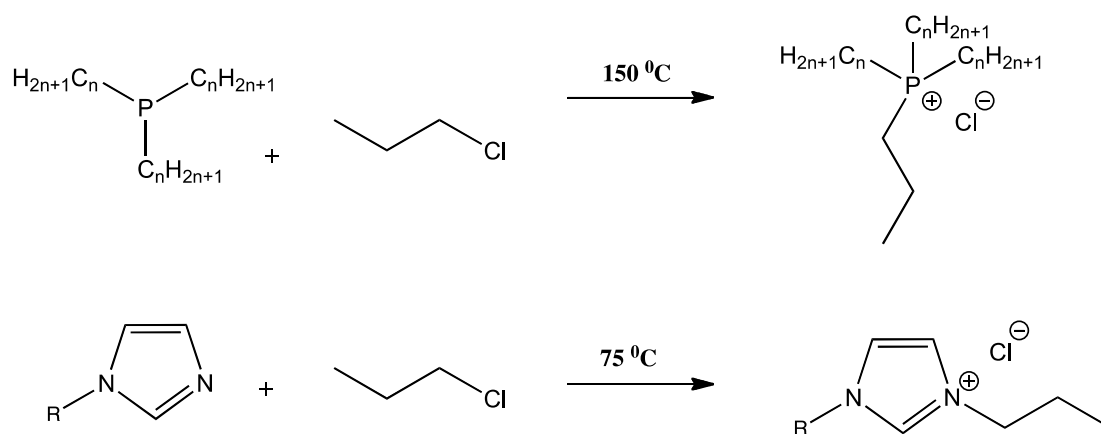
These unique properties are of benefit for chemical sensing applications, as a good knowledge on the overall chemistry of the IL can be used to improve on the limitations of previous sensing approaches.

2.1. Synthesis of ILs

The current literature is dominated by two main cation-families, ammonium and phosphonium based ILs. The quaternisation of phosphines has been described in the literature by Bradaric *et al.*[27] and Ermolaev *et al.*[28], whilst the corresponding reaction for amines is described by Busi *et al.*[29]. The reaction schemes are somewhat similar in that both electron rich starting materials undergo a S_N2

addition reaction in the presence of a haloalkane at elevated temperatures over time. A general reaction scheme for the quaternisation of a trialkylphosphine with 1-chloropropane to form the phosphonium salt is presented in figure 1 (top); whilst the quaternisation of 1-substituted imidazole with 1-chloropropane to form the imidazolium salt is shown in figure 1 (bottom). Both reactions can be undertaken without solvent, as the reactants are liquid in both cases. Once the reaction is complete, the new IL formed is easily isolated by vacuum removal of the volatile halo-alkanes.

Figure 1. Direct nucleophilic addition of a trialkylphosphine (**top**) and 1-substituted imidazole (**bottom**) with 1-chloropropane to form their corresponding chloride ILs.

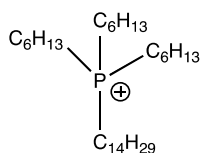


Variants of the halide ILs can be prepared via their ion-exchange metathesis reaction with group I organic anions[30,31]. Key to the success of this reaction is the choice of solvent that it is undertaken in, it must serve to solvate the new IL formed and *preferably* promote the precipitation of the by-product (a classical alkali-halide salt). As the IL begins to form over time, so too do the ionic interactions of alkali and halide atoms, respectively, which transfer out of the reaction solvent. The new IL can then be isolated by filtration of the by-products and drying over time.

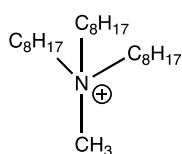
A list of the of the more popular ions produced via quaternisation and ion-exchange reactions can be seen in figure 2; beginning with the tetra-alkylated phosphoniums and ammoniums (i) and (ii), N-heterocyclic amines (iii) and (iv), the amides (v) and (vi), hexafluorinated phosphates (vii) and (viii), benzenesulfonates (ix) and (x) and finally some group III tetrahalides (xi) and (xii).

Figure 2. A list of the more popular ions used throughout the literature for IL syntheses. Cations are listed on the left ((i) – (iv)), and anions are listed on the right ((v) – (xii)).

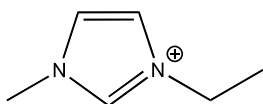
Cations



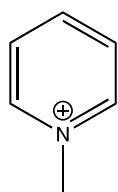
(i) [P_{6,6,6,14}]⁺



(ii) [N_{1,8,8,8}]⁺

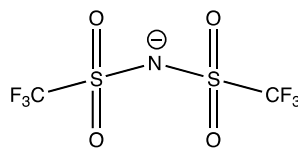


(iii) [C₂mIm]⁺

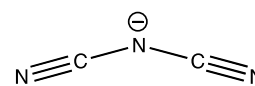


(iv) [mpyr]⁺

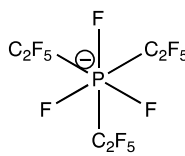
Anions



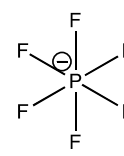
(v) [NTf₂]⁻



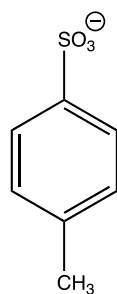
(vi) [DCA]⁻



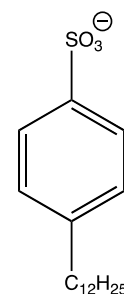
(vii) [FAP]⁻



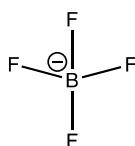
(viii) [PF₆]⁻



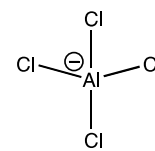
(ix) [TOS]⁻



(x) [DBSA]⁻



(xi) [BF₄]⁻



(xii) [AlCl₄]⁻

3. Stimuli Responsive Materials

Stimuli responsive materials (SRM's) are those that can undergo (in some cases a reversible) change in their molecular configuration in response to an external applied stimulus. They can of course be subdivided in relation to the particular stimulus applied, and are the subject of previous reviews. These reviews have discussed the alteration of molecular configurations in response to irradiated light[32], to an applied voltage[33], to a change in temperature[34], to an incident magnetic field[35] or a change in the pH of the surrounding environment[36]. The change in molecular configuration is usually accompanied by an observable signal, for example a change in colour, conductivity or surface energy. They are therefore ideal candidates for use as chemical sensors; if the molecular rearrangement can be induced by an interaction with a defined analyte, then it can also be used as the chemical sensor signal. Photo responsive materials are particularly good candidates for chemical sensors as, in some cases, the incident irradiation can be low enough in power (such as in the use of low power LEDs [37,38]) to be non-invasive on the sensing materials. This is an important consideration for materials that may be subject to photostability issues, as exposure to high power sources can lead to rapid decomposition of the material. Addressing this issue can lead to improved reproducibility in the response obtained, which will improve the performance of the sensing device over increased time periods[39].

3.1 Photo-responsive materials

The *cis-trans* isomerization of azo compounds and stilbene has also been used to produce photo-responsive solutes for switching the viscosity of the resulting solution [40,41]. Azobenzenes are particularly good candidates for use as chemical sensors as the photoisomerisation event not only yields a change in colour[42], but also a significant change in polarity [43,44]. The use of organogelators (a molecule which exhibits significant electrostatic interactions leading to the formation of an interconnected network[45,46]) has proved a worthy route for the development and incorporation of the photo-responsive chemistries of azo compounds into the solid state. Murata *et al.*[47] developed a cholesterol-based gelator containing azobenzene; irradiation of the system at 330 - 380 nm led to isomerization of the *trans* azo linkage to its *cis* conformation and thereby disrupting gelation. Irradiation of the sol at wavelengths longer than 460 nm resulted in reversal of the isomerization and re-formation of the gel. Pozzo and co-workers[48] employed the photochromic equilibrium of 3,3-diphenyl-3*H*-naphthopyran, which brings about a large conformational change. Two forms of the naphthopyran can be distinguished: a coloured open form and a colourless closed form[48]. In the closed form the naphthopyran unit does not influence the stacking process and the molecule acts an efficient gelator. Irradiation with UV light converts the naphthopyran unit into the open form and prevents the carbamate group from stacking. The isomerization is accompanied by a colour change, and while the gel liquefies it becomes yellow. Heating converts the pyran unit back into the closed form, and upon cooling a colourless gel is again obtained (Figure 3).

Of particular interest to many are the photo-responsive molecule Spiropyran (SP) and its zwitterionic isomer Merocyanine (MC)[49,50]. The two isomers exist in a photo-dynamic equilibrium controlled by the application of UV and visible light, respectively [51,52]. Furthermore, MC can act as a Lewis base in the presence of an acid or metal ion solution, forming protonated forms and ion-complexes. Particularly interesting is that the ion/proton binding can be reversed using white light, releasing the bound species and reforming the inactive SP isomer. In addition to the SP and MC isomers, the protonated and metal-ion chelated forms exhibit unique optical properties and the entire system is therefore inherently self-indicating of their status (Figure 4).

Figure 3. Light-induced switching between open and closed isomers of naphthopyrans[48].

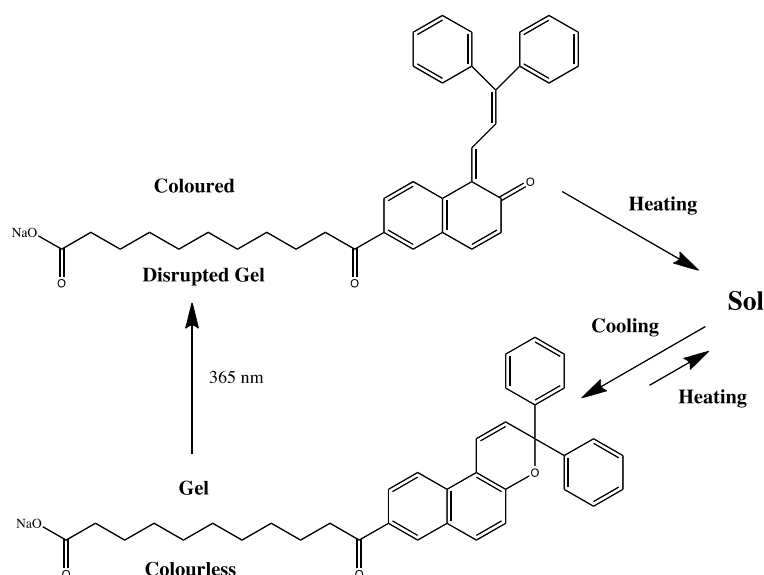
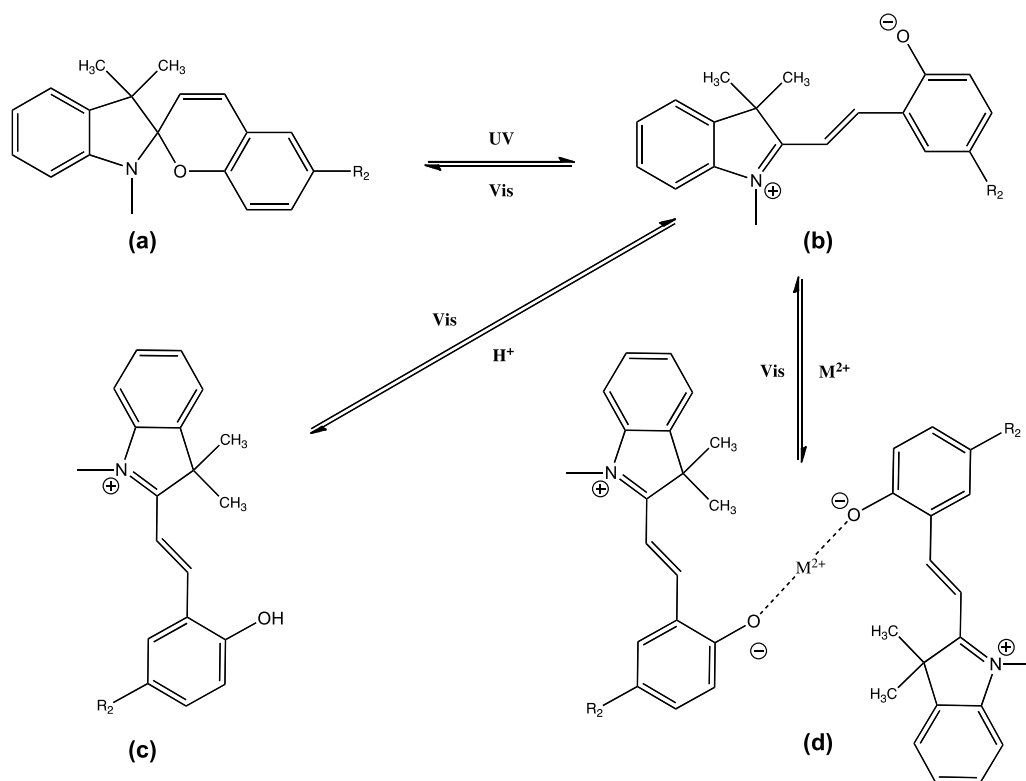


Figure 4. The photodynamic, pH and metal-ion chemistry of SP materials; (a) SP, (b) MC, (c) MC-H and (d) MC-M²⁺ complex.



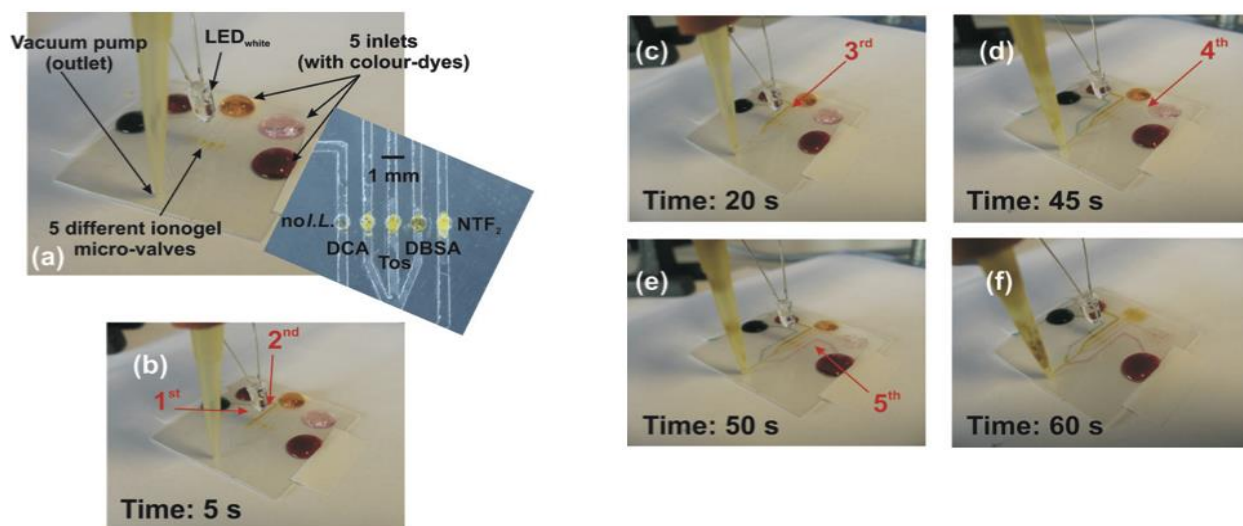
Previous research has explored the behaviour of the SP and MC isomers in systems capable of user controlled transition metal ion uptake and release[39,53], as effective solvatochromic probes[54,55] and as hybrid materials that exhibit user controlled multi-switchable optical properties[56]. Combining the photo-chemistry of SP and MC with the responsive chemistries of a hydrogel was first investigated by Szilagy *et al.*[57]. We have expanded the area slightly to combine the properties of photo-responsive hydrogels with ILs producing ionogels[18,58], which will now form the basis of the next discussion.

4. Stimuli Responsive Ionogels

4.1 Photo responsive ionogels for direct fluid control in microfluidic devices

One of the issues for microfluidic or “lab on a chip” devices is the controlled movement of liquid throughout the device. The practical solution has been to employ high power consuming pumps and/or actuating valves, which mean that most prototype devices are anything but a “lab on a chip”. One novel alternative to this approach, is to control fluid movement using stimulus-responsive polymer valves integrated into the fluidic system that can be controlled using light [18]. The valve was based on an ionogel of which there were two distinct components: **(a)** The polymer gel, a co-polymer based on (poly(*n*-isopropylacrylamide) (pNIPAAm) and SP, and **(b)** phosphonium ILs.

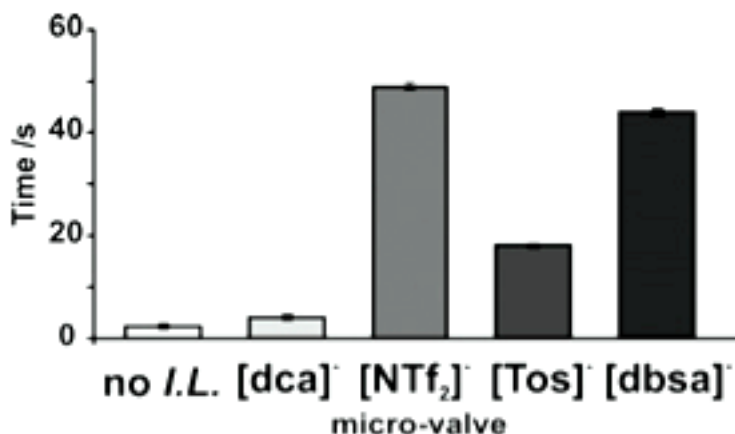
Figure 5. Overview of the photo-responsive ionogels: (a) micro-valves closed; the applied vacuum is unable to pull the dyes through the micro- channels, (b) white light is applied for the first time. (c) – (f) controlled fluidic flow as a function of the IL anion in the ionogel [18] – *Reproduced by permission of The Royal Society of Chemistry, <http://dx.doi.org/10.1039/B914709H>.*



pNIPAAm is a well known hydrogel that can swell and contract thereby drawing in and expelling aqueous liquids as a function of pH and temperature[59,60]. In this work, the authors initially soaked the ionogels in acidic solution, which caused them to swell, and protonated the SP as part of the polymer backbone (Figure 4 (b) -> (c)). Figure 5 is a brief summary of the work, in which the ionogels were incorporated into a microfluidic device based on poly(methylmethacrylate) (PMMA). Four ionogel devices were prepared based on phosphonium cations with amide and benzenesulfonate anions.

Irradiation of the ionogels with white light caused the protonated MC isomer to revert back to its SP closed form. This induced a change in the polarity of the ionogel (hydrophilic, MC-H⁺ to hydrophobic, SP), and the subsequent expulsion of water from the gel, causing it to contract. The role of the IL proved crucial in controlling the rate of contraction (and hence the fluid movement) as the thermal relaxation of the MC-H⁺ to SP is significantly affected by its physico-chemical interactions with the IL[54,55]. Through variation of the ionogel composition; the micro-valves were tuned to open at different rates when illuminated under the same white light source (Figure 2 (c) – (f), and figure 6.)

Figure 6. Actuation times of photo-responsive ionogel valves under uniform white light illumination as a function of the IL anion[18] - Reproduced by permission of The Royal Society of Chemistry, <http://dx.doi.org/10.1039/B914709H>.



4.2 Electro responsive ionogels for sensing applications

A particular advantage of using ionogels occurs in applications where the intrinsic ionic conductivity of the ionogel can be exploited, for example in electrochemical sensors and devices. Ion-selective electrodes (ISE's) are a particular breed of electrochemical sensor, which convert the activity of a given analyte in solution into a voltage potential. ISE's employ polymer gels based mainly on poly(vinylchloride) (PVC) and PMMA[61,62], for the selective detection of important environmental and biological analytes at levels as low as nanomolar concentrations in some cases[63,64]. As PVC and PMMA exhibit high glass transitions, organic plasticisers are used to produce flexible transparent polymer membranes. They are prepared by co-dissolution with a suitable molecular solvent, producing the membrane as the solvent evaporates[65]. The sensing agents are often incorporated into the polymer membranes via their co-dissolution with the polymer and the plasticiser solution.

A prerequisite for a plasticiser therefore is that it must exhibit a markedly lower glass transition itself. Phosphonium ILs have been shown to exhibit glass transitions as low as -77°C [66], and have been shown to plasticise both PMMA and PVC, producing flexible films suitable for use in ISEs [67,68].

A particular advantage of using ILs as plasticisers for PVC/PMMA gels is that they can further replace some of components needed for a functioning ISE. For example, an ion-exchanging salt is typically used in ISE membranes to facilitate the movement of the analyte between the aqueous and polymeric phases[69]. Chernyshov *et al.* have reported the selective detection of dopamine and adrenaline at concentrations as low as 1×10^{-7} M, using an IL as plasticiser without the need for an ion-exchanging salt[70].

Ionogels have also recently found use in amperometric sensors. Nádherná *et al.* described the use of an ionogel based on poly(ethyleneglycol) and 1-butyl-3-methylimidazolium hexafluorophosphate [C₄mim][PF₆] as an amperometric sensor for nitrogen dioxide. Using the ionogel as the electrolyte, the authors reported a detection limit as low as 0.3 ppm[71].

4.3 Electro responsive ionogels for electrochromic applications

Electrochromic materials undergo a reversible change in their optical properties in response to an applied voltage[72]. They are good candidates for sensing templates as the colour can vary by the choice of the electrochrome (of which there are many[73,74]) and its concentration. The colour generated can therefore act as a sensor signal in response to an electrical potential[75]. Ionogels are good candidates for electrochromic devices as their intrinsic ionic conductivity can facilitate the charge required to produce the optical response[17].

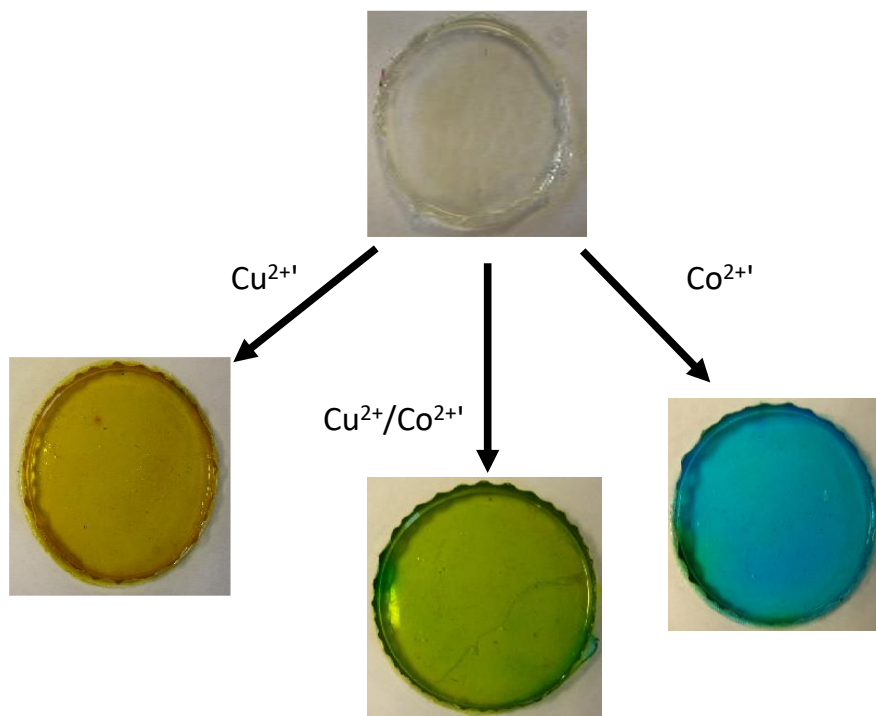
A publication by Ahmad *et al.* describes the development of a complementary electrochromic device based on Tungsten oxide and Prussian blue[76]. The authors prepared an ionogel using sol-gel hydrolysis chemistry of Tetraethyl orthosilicate (TEOS) in the presence of the IL 1-ethyl-3-methylimidazolium bis(trifluoromethane)sulfonamide [C₂mIm][NTf₂], which formed a solid-state device upon solidification between two electrodes. In this case the responsive chemistry of the ionogel facilitated the current flow toward the electrodes (to which the chromophores were deposited), which in turn caused a change in the optical properties of the device. The authors reported an ionic conductivity of 1.2×10^{-2} S/cm for the ionogel, which allowed for fast switching times whilst the colouration efficiency of the device was calculated to be 64 cm²/C.

Recent publications have detailed the use of ionogels as the solid-state electrolyte for electrochromic devices based on 4,4' – disubstituted bipyridinium salts (viologens)[17,77], poly(aniline)[78], poly(pyrrole)[79] and Prussian Blue[80].

4.4 Optically responsive ionogels for sensing applications

Polymer optodes are similar to ISE's in terms of their composition, and how the analyte transfer is facilitated from the sample into the membrane phase[81]. Ionogels have been employed within optodes, mainly as the plasticiser to produce the film. Zhu *et al.* used phosphonium based ionogels with a conventional chromoionophore for the detection of important inorganic acids such as HCl and H₂SO₄[82]. The authors detailed enhanced selectivity toward hydrophilic anions for the ionogels versus conventional plasticisers, which they have attributed to the increased dielectric constant of the ionogel. A publication based on a similar concept by Topal *et al.* detailed the increased selectivity of zinc phthalocyanines to pH when incorporated into imidazolium ionogels[83]. One particular publication on ionogels as optodes detailed significant template simplification (as a result of the particular IL used), whilst the ionogel/optode produced three distinct colours in the presence of Cu²⁺ and Co²⁺ ions[15] as shown in Figure 7.

Figure 7. Optical response of *trihexyltetradecylphosphonium dicyanamide* ($[P_{6,6,6,14}][DCA]$)/PVC based ionogel upon complexation with Cu^{2+} (yellow, left), Co^{2+} (blue, right), and a mixture of the two (green, middle)[15].



Lunstroot *et al.* have pioneered the concept of ionogels as solid-state luminescence devices in two separate publications[84,85]. In both cases the ionogels were based on a siloxane support, whilst the IL was used to bind to lanthanide element that exhibited photoluminescence upon UV irradiation. In fact Cheminet *et al.* recently expanded the concept of optically responsive ionogels to fluorescent materials[86]. A phenylene–ethynylene oligomer was synthesised and chemically tethered to an imidazolium cation; which exhibited fluorescence in the solid state as part of a siloxane based ionogel.

4.5 Ionogels as bio-sensing components

Along with the electrochemical applications of ILs (as discussed), ILs have gained momentum in bio applications. Recent work in the area include ILs as biocatalytic reactions[87,88], biosensors[89], protein stabilization[90] and biopreservation[91]. They have been proposed as unique solvents for biomolecules such as proteins / enzymes because of their unusual solvation characteristics. Work by Fujita *et al.*[89,92] and others[90,93] have shown that some proteins are, in fact, soluble, stable and remain active in some ILs. As a case study Cytochrome c (cyt. c) was found to have enhanced solubility and stability in a biocompatible IL solution based on the dihydrogen phosphate anion[92]. This is an important observation since proteins are sometimes unstable when handled *in vitro*, and stabilizing agents are a necessary component to ensure their long-term stability.

This is especially true of proteins that have pharmaceutical potential since lack of stability is a limitation to widespread use of some protein therapeutics. It has been well documented that enzyme performance in an IL is affected by several parameters including water activity, pH and impurities[94].

Other important factors that play a role in enzyme stability / activity include IL polarity, hydrogen bond basicity and nucleophilicity of anions, ion kosmotropicity and viscosity.

Although outside the scope of this discussion, these areas have been discussed in an excellent review by Zhao[95]. Abe *et. al*[96] recently synthesized a number of phosphonium salts that have an alkyl ether group present. The phosphonium salts moiety is commonly found in living creatures, and it was hypothesized that this family of ILs have good affinity with enzyme proteins and may provide a good environment for enzymes. Some examples are based on the immobilization of enzyme-IL systems in chitosan[97] or Nafions[98]. Thus, catalytically active proteins and enzymes may also be confined for biosensors applications in order to achieve direct electron transfer within ionogels.

It is therefore proposed that incorporating these biocompatible ILs into ionogels is a particularly attractive strategy in the field of biosensing. These materials, in theory, will inherit all of the favourable IL properties whilst being in a solid, gel like structure[99].

4.5.1 The Need For Wearable Sensors

Wearable sensors allow the continuous monitoring of a person's physiology in a natural setting. At present, health-monitoring systems using electronic textiles are mainly targeting applications based upon physiological parameter measurements, such as body movements or electrocardiography (ECG). However, due to their relative complexity, there is very little activity in the development of real-time wearable chemo/bio sensing for sports applications.

Micro-Total-Analysis-Systems (μ TAS) and Lab-on-a-Chip (LOAC) technology are widely used in analytical chemistry and biotechnology[100] but they are still rarely used in other areas like sports science. In this field, wearable sensors are becoming increasingly employed, through the use of embedded transducers or smart fabrics for monitoring parameters like breathing rate, heart rate and footfall[101]. These sensors require that the desired sample of analysis, usually a body fluid such as sweat is delivered to the sensor's active surface, whereupon a reaction happens and a signal is generated. Ideally, the system must be low cost, while still being robust, miniature, flexible, washable, reusable or disposable[1]. All these requirements point to micro-fluidic devices as the key tools for improving wearable chemo-/bio-sensing[102].

There are several factors that correlate sweat pH and health. Changes in the pH of the skin are reported to play a role in the pathogenesis of skin diseases like irritant contact dermatitis and acne, among others.[103] Patterson *et al.*[104] showed that inducing metabolic alkalosis through the ingestion of sodium bicarbonate led to increased blood and sweat pH. Furthermore, it has been reported that sweat pH will rise in response to an increased sweat rate.[105] A relationship was also observed between pH and sodium (Na^+) levels in isolated sweat glands in that the greater the concentration of Na^+ , the higher the sweat pH will be[106]. As exercising in a dehydrated condition has been shown to lead to increased levels of Na^+ , it can be seen that such changes can be measured directly (e.g. using a Na^+ selective sensor) or indirectly by monitoring sweat pH.[107]

In order to provide sensors with good sensitivity, selectivity and stability, various support materials, methods and reagents for immobilization of pH indicators were employed by our group[108,109]. . . A barcode system (as shown in figure 8) was developed as an initial sweat sensor with an ionogel being an important component in the fabrication of the sensing platform[110].

Figure 8. Barcode system for the analysis of human sweat developed by Benito-Lopez *et al.*[110]. Ionogels were based on pNIPAAm and Phosphonium ILs, whilst the change in optical properties of the encapsulated colorimetric dye was used to correlate with the pH of the incident sample.



The ionogel consisted of two monomeric units; (NIPAAm) and N,N-methylene- bis(acrylamide) (MBAAm) in the ratio 100:5, respectively[57]. The reaction mixture was prepared by dissolving both monomeric sub-units and the photo-initiator dimethoxy-phenylacetophenone (DMPA) into the IL([P_{6,6,6,14}][DCA]). The barcode (18 x 10 mm), consisted of four independent reservoirs, and was easily fabricated in PMMA and pressure-sensitive adhesive in five layers using CO₂ ablation laser[110]. Immersing the barcode in de-ionised water, and then varying the pH from 0 to 14 in intervals of one pH unit allowed the stability of the barcode to be studied. The pH was monitored using a benchtop pH meter and it was repeated at least four times without observing any damage to the barcode and/or ionogel.

The ionogel matrix proved to be very robust even at harsh pH conditions (0-14) and it was shown that the pH indicators bromocresol green, bromocresol purple and bromothymol blue retained their pH indicator properties within the ionogel. The ionogel-dye interactions ensure no leaching of the dyes occurs during experiments, thereby providing long durability of the device for the monitoring of sweat pH measurements over time[110,111].

Important contributions to current sensor research in the area of LOAC or μ TAS systems can be made by employing ionogels as active constituents within organic electrochemical transistors (OECTs). Many of the applications for OECT / ionogel sensors will likely involve disposable devices, cheap fabrication and therefore enzyme stability is of the utmost importance. OECT devices are inherently low-power and relatively easy to fabricate.

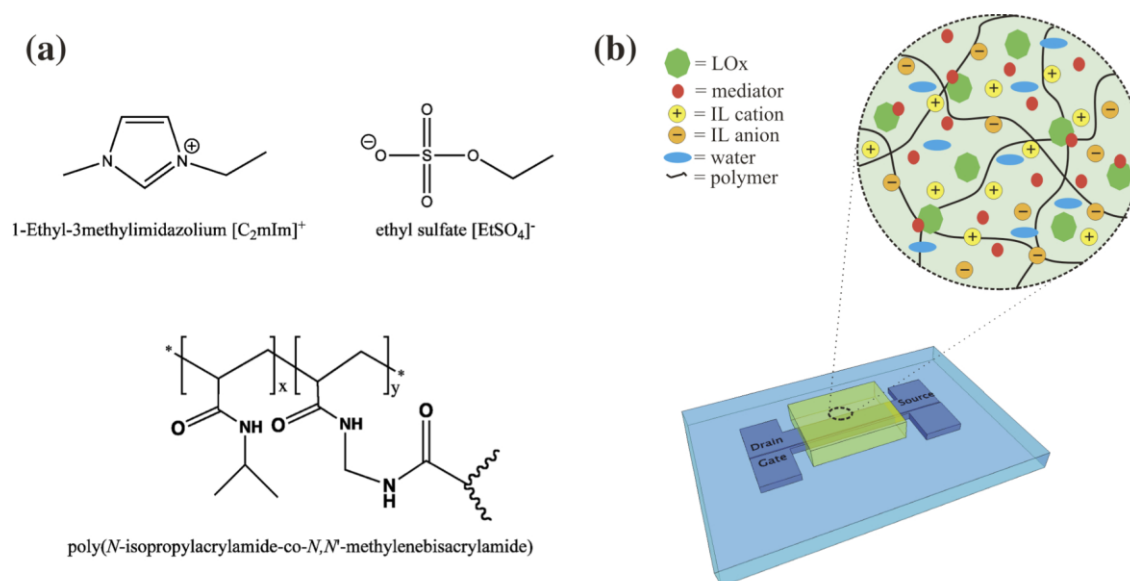
Nilsson *et al.*[112] demonstrated the potential for OECT sensors to be manufactured inexpensively using printing techniques for mechanically flexible single-use tags. They fabricated humidity sensors on thin polyester foils and on paper. Combining OECT properties with those of ionogels therefore

offers significant potential for realizing new generations of solid-state biosensing devices in a variety of form factors, and using ILs to optimise the stability and reactive nature of the host enzyme.

Most recently the development of a solid-state electrolyte incorporated into an OECT has been reported for the detection of lactate[20]. The authors report the possibility of a fast, reliable, robust, and cheap way of measuring lactate concentration in physiological fluids using an ionogel as the basis of a biosensor (Figure 9) which will open the way to lactate biosensors for health and sport applications. Current methods of detection of lactate include conducting polyaniline films[113], carbon nanotubes[114], screen printed Prussian blue electrodes[115], and biosensors based on electrochemiluminescent detection[116]. Commercial lactate sensors are also available,[117] based on standard electrochemical methods. One example is the lactate SCOUT (Senslab), which however, performs measurements directly in blood, making real-time detection invasive and impractical for long-term studies.

The detection of lactate (deprotonated form of lactic acid) in blood provides a biochemical indicator of anaerobic metabolism in patients with circulatory failure[118]. In addition to its presence in blood, lactate can be found in sweat (concentration range between 9 to 23 mM), reflecting, in an indirect way, eccrine gland metabolism[119]. It is well known that lactate concentration increases during physical exercise, making it a useful parameter to monitor wellness, physical fitness and the effects of exercise[120].

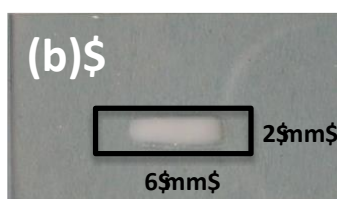
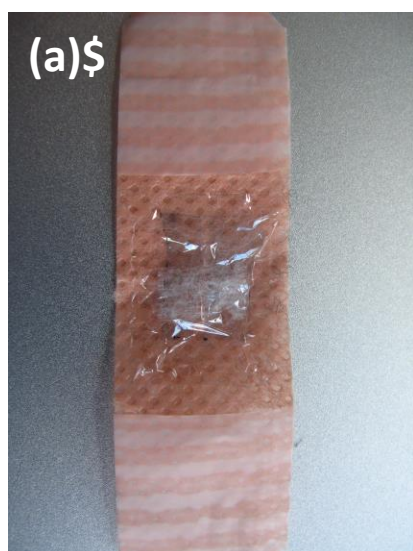
Figure 9. (a) Ionogel components and (b) a schematic representation of the OECT device with ionogel / enzyme mixture[20]. *Reproduced by permission of The Royal Society of Chemistry, <http://dx.doi.org/10.1039/C2JM15716K>.*



The possibility to pattern PEDOT:PSS in a wide variety of substrates such as glass, flexible plastic and textiles, opens new routes for the development of wearable biosensors. Such biosensors can therefore be incorporated into fabrics such as t-shirts, sweat bands or shorts allowing for the analysis of real time measurements of the target biomolecules. Figure 10, (left) shows a proof of concept of this type prepared in our laboratories which consists of an OECT transistor incorporated into a plaster,

while (right) shows an ionogel formulation printed using an inkjet printer (volume in the nano liter range). However, more studies are needed to understand device degradation mechanisms and improve sensor lifetimes.

Figure 10. A printable OECT / Ionogel composition incorporated into a widely available fabric (plaster) and the development of printable ionogels. Volumes are in the nL range.



- 15 ink' jet layers.
- 90 seconds UV photo' polymerisa; on.
- 150 μm layer thickness.

5. Whats does the future of stimuli responsive ILs hold?

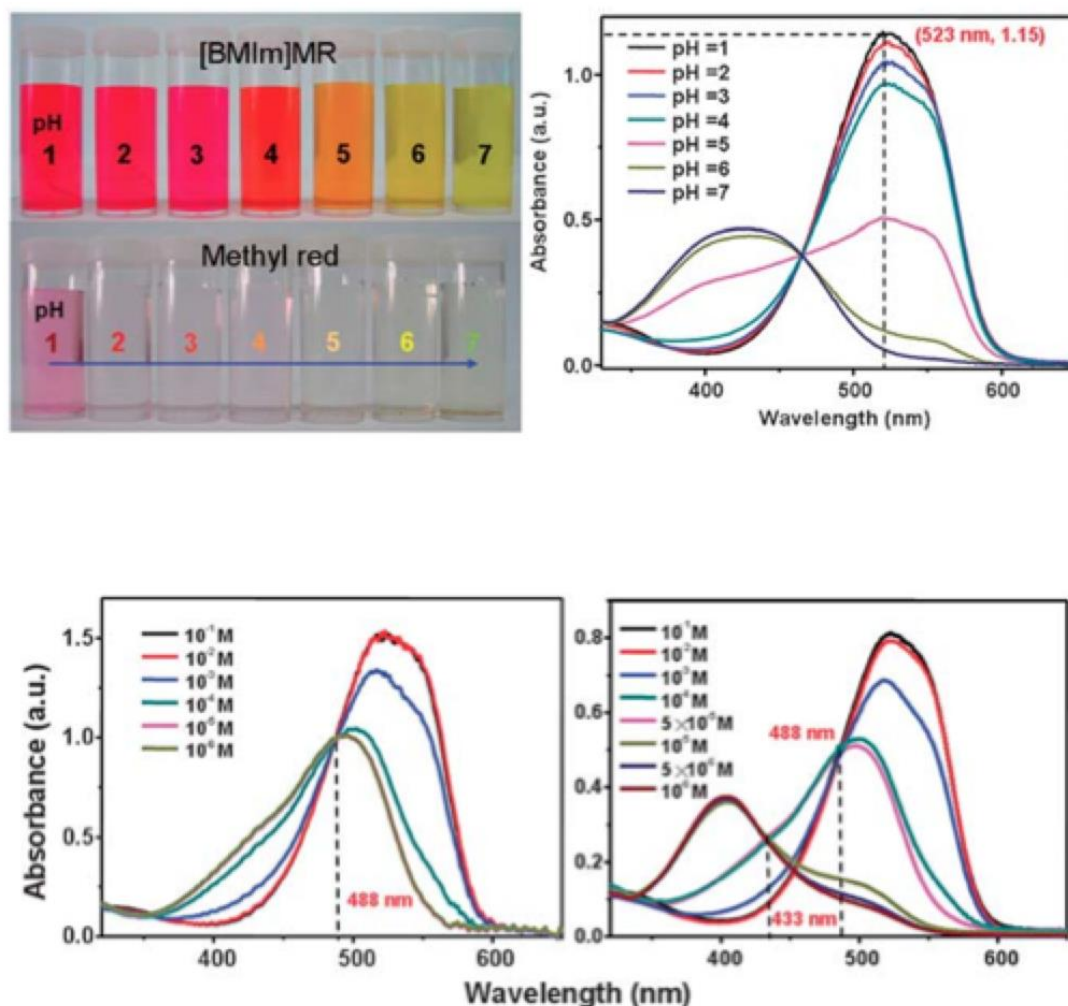
5.1 pH, photo-responsive ILs

As the explosion of interest in ILs as diverse, functional agents for multidisciplinary research continues, so too has the advent of ILs exhibiting stimuli responsive chemistries. This represents a subsection topic for both respective fields of research (ILs and SRM's). Previously solid, lattice structured and anionic SRMs have been shown to form liquids when their respective counter ions are exchanged for those used in IL studies (Figure 2 (i) – (iv)). The new materials therefore exhibit their previous responsive chemistries in the liquid phase, combined with properties inherent to ILs (negligible vapour pressure, ionic conductivity etc.). Recently, two individual communications described ILs based on the photo and pH responsive chemistries of azobenzene.

Pina *et al.* first described the synthesis of ILs based on the azobenzene dye Methyl Orange (MO)[121]. ILs were prepared by allowing the sodium salt of MO to undergo ion-exchange metathesis with halide ILs of imidazolium, phosphonium, sulfonium and guanidinium respectively, yielding intrinsically photochromic ionic liquids. The paper details how the behaviour of photochromic ILs can be tuned by choice of the cation, as measured by the individual thermal relaxation rates from the *cis* back to *trans* structure of four different ILs dissolved in water and ethanol.

Zhang *et al.* further developed on the concept of azobenzene functionalised ILs in a separate publication. In this paper, ILs based on imidazolium and pyrrolidinium cations and methyl red (MR) were prepared for use as acido-responsive sensors in aqueous and non-aqueous solutions[122]. MR differs slightly to MO in that a carboxylate group replaces the sulfonate in position R₁. The authors detail the advantages in preparing ILs with MR as shown in figure 11.

Figure 11: (Top) Increased selectivity of [C₄mIm][MR] over the conventional Na⁺ salt of MR from pH 1 to 7, and (bottom) the increased selectivity of the [C₄mIm][MR] (right) to HOTf over the conventional sodium salt in ethanol (left)[122] - Reproduced by permission of The Royal Society of Chemistry, <http://dx.doi.org/10.1039/C0AN00885K>.



The sodium salt of MR was shown to be sparingly soluble in water, meaning its pH responsive chemistry is void in this solvent. By exchanging sodium for [C₄mIm]⁺, the resultant IL was completely miscible with water and the pH responsiveness of the chromophore allowed a selective response to be observed in pH's ranging from 1 – 7 (figure 11 (top)).

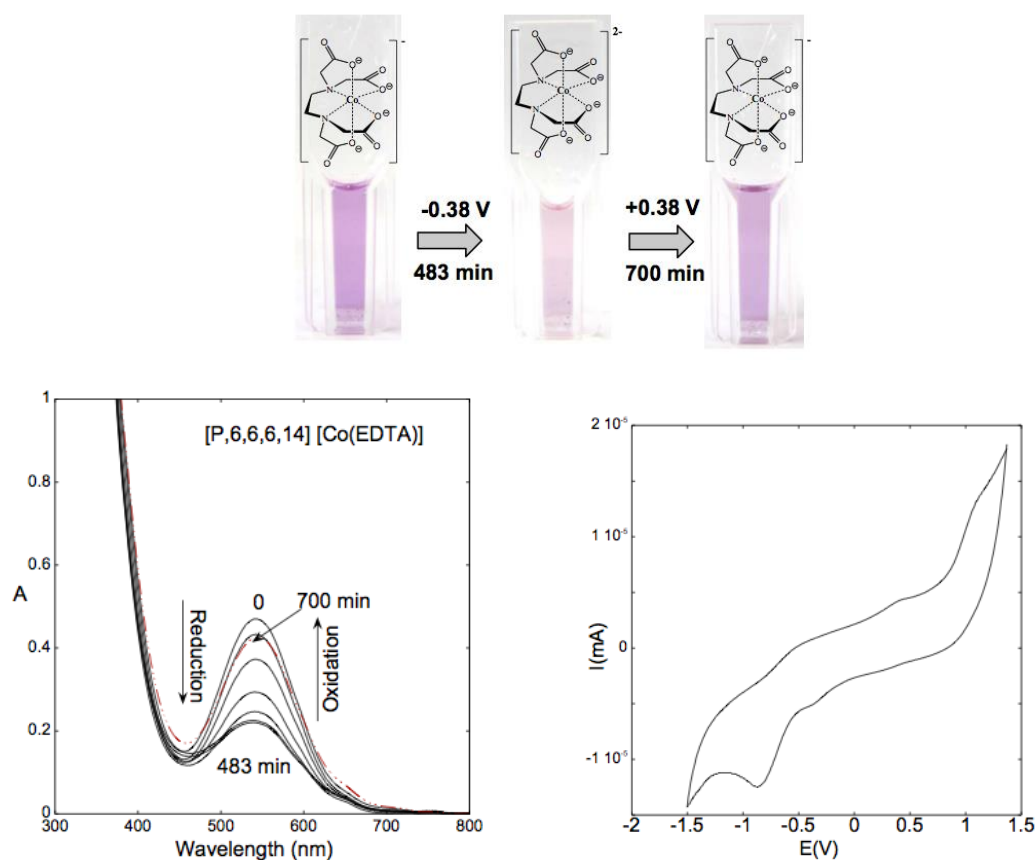
Furthermore, an increased selectivity for the same IL over MR for trifluoromethanesulfonic acid (HOTf) was observed in solutions of ethanol (figure 11, (bottom)). The authors postulate a differing response mechanism for the ILs response over MR and attribute it to the increased response to HOTf at low concentrations.

This particular publication highlights the novelty and advantages of preparing ILs with stimuli responsive chromophores. The unique ionic nature of these liquids allows well-known chemistries to be performed in previously inhospitable solvent environments.

5.2 Electrochromic ILs

A recent communication by Pina *et al.* described the syntheses of intrinsically electrochromic ILs based on the transition metal complexes of ethylenediaminetetraacetic acid (EDTA)[123]. Again the ILs were prepared via ion-exchange of its Na^+ salt with the relevant halide salt of cations based on phosphonium, imidazolium and ammonium derivatives. The spectroelectrochemical response obtained for the phosphonium IL is shown below in figure 12.

Figure 12: (Top) The reversible optical redox chemistry of $[\text{P}_{6,6,6,14}][\text{Co}(\text{EDTA})]$. (Bottom, Left) Change in absorbance as a function of the applied voltage and (Bottom, Right) the CV obtained for of $[\text{P}_{6,6,6,14}][\text{Co}(\text{EDTA})]$ [123] - Reproduced by permission of The Royal Society of Chemistry, <http://dx.doi.org/10.1039/C0CC03892J>.



In this case the electrochromic functionality of the IL is a result of the coordinated central atom switching between two distinct redox states ($\text{Co}^{2+/3+}$). A secondary electrolyte was not needed as the intrinsic conductivity of the IL facilitated the current between two electrodes.

6. Concluding remarks.

Our overview has attempted to present the current state of development in the field of employing stimuli responsive ionogels as active components in chemical sensors, as reported in the open literature. Arguably the most advantageous feature of using ionogels within sensing templates is the wide degree of ion choice, and related breadth of tuneability of the properties of the IL and therefore the behaviour of ionogels that are based on them (be it either the cation or anion of the IL, or the chemical structure of the repeating unit). We have highlighted the applications of ionogel chemistry in diverse research areas such as photo-controlled valves in microfluidic devices, as active components in optical/electrochemical sensors and in the development of robust bio-chemical sensing platforms. We have also highlight a new emerging area in the development of stimuli responsive ILs, where the responsive chemistries of existing materials are integrated into an IL in the liquid phase.

The use of ionogels in sensing platforms clearly has several advantages over current technologies. There is great potential in the development of ionogel science, with a diverse area of application ready to be explored.

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References and Notes

1. Diamond, D.; Coyle, S.; Scarmagnani, S.; Hayes, J. Wireless Sensor Networks and Chemo-/Biosensing. *Chem. Rev.* **2008**, *108*, 652-679.
2. Bobacka, J.; Ivaska, A.; Lewenstam, A. Potentiometric ion sensors based on conducting polymers. *Electroanalysis* **2003**, *15*, 366-374.
3. Singh, N.; Mulrooney, R.C.; Kaur, N.; Callan, J.F. A nanoparticle based chromogenic chemosensor for the simultaneous detection of multiple analytes. *Chemical Communications* **2008**, 4900-4902.
4. Schmittel, M.; Lin, H.W. Quadruple-channel sensing: A molecular sensor with a single type of receptor site for selective and quantitative multi-ion analysis. *Angewandte Chemie-International Edition* **2007**, *46*, 893-896.
5. Bakker, E.; Diamond, D.; Lewenstam, A.; Pretsch, E. Ion sensors: current limits and new trends. *Analytica Chimica Acta* **1999**, *393*, 11-18.
6. Diamond, D. Internet-scale sensing. *Analytical Chemistry* **2004**, *76*, 278A-286A.
7. Park, C.O.; Fergus, J.W.; Miura, N.; Park, J.; Choi, A. Solid-state electrochemical gas sensors. *Ionics* **2009**, *15*, 261-284.

8. Anastasova-Ivanova, S.; Mattinen, U.; Radu, A.; Bobacka, J.; Lewenstam, A.; Migdalski, J.; Danielewskic, M.; Diamond, D. Development of miniature all-solid-state potentiometric sensing system. *Sensors and Actuators B-Chemical* **2010**, *146*, 199-205.
9. Ahn, S.-K.; Kasi, R.M.; Kim, S.-C.; Sharma, N.; Zhou, Y. Stimuli-responsive polymer gels. *Soft Matter* **2008**, *4*, 1151-1157.
10. Osada, Y.; Gong, J.P. Soft and wet materials: Polymer gels. *Advanced Materials* **1998**, *10*, 827-837.
11. Neouze, M.A.; Le Bideau, J.; Gaveau, P.; Bellayer, S.; Vioux, A. Ionogels, new materials arising from the confinement of ionic liquids within silica-derived networks. *Chemistry of Materials* **2006**, *18*, 3931-3936.
12. Susan, M.A.; Kaneko, T.; Noda, A.; Watanabe, M. Ion gels prepared by in situ radical polymerization of vinyl monomers in an ionic liquid and their characterization as polymer electrolytes. *Journal of the American Chemical Society* **2005**, *127*, 4976-4983.
13. Le Bideau, J.; Viau, L.; Vioux, A. Ionogels, ionic liquid based hybrid materials. *Chemical Society Reviews* **2011**, *40*, 907-925.
14. Ueki, T.; Watanabe, M. Macromolecules in ionic liquids: Progress, challenges, and opportunities. *Macromolecules* **2008**, *41*, 3739-3749.
15. Kavanagh, A.; Byrne, R.; Diamond, D.; Radu, A. A two-component polymeric optode membrane based on a multifunctional ionic liquid. *Analyst* **2011**, *136*, 348-353.
16. Kavanagh, A.; Hilder, M.; Clark, N.; Radu, A.; Diamond, D. Wireless radio frequency detection of greatly simplified polymeric membranes based on a multifunctional ionic liquid. *Electrochimica Acta* **2011**, *56*, 8947-8953.
17. Kavanagh, A.; Copperwhite, R.; Oubaha, M.; Owens, J.; McDonagh, C.; Diamond, D.; Byrne, R. Photo-patternable hybrid ionogels for electrochromic applications. *Journal of Materials Chemistry* **2011**, *21*, 8687-8693.
18. Benito-Lopez, F.; Byrne, R.; Raduta, A.M.; Vrana, N.E.; McGuinness, G.; Diamond, D. Ionogel-based light-actuated valves for controlling liquid flow in micro-fluidic manifolds. *Lab on a Chip* **2010**, *10*, 195-201.
19. Cicmil, D.; Anastasova, S.; Kavanagh, A.; Diamond, D.; Mattinen, U.; Bobacka, J.; Lewenstam, A.; Radu, A. Ionic Liquid-Based, Liquid-Junction-Free Reference Electrode. *Electroanalysis* **2011**, *23*, 1881-1890.
20. Khodagholy, D.; Curto, V.F.; Fraser, K.J.; Gurfinkel, M.; Byrne, R.; Diamond, D.; Benito-Lopez, F.; Owens, R.M. Organic electrochemical transistor incorporating an ionogel as solid state electrolyte for lactate sensing. *J. Mater. Chem* **2011**, *Accepted Dec 2011*.
21. Welton, T. Room-temperature ionic liquids. Solvents for synthesis and catalysis. *Chem. Rev.* **1999**, *99*, 2071-2083.
22. Fraser, K.J.; MacFarlane, D.R. Phosphonium-Based Ionic Liquids: An Overview. *Australian Journal of Chemistry* **2009**, *62*, 309-321.
23. Zhao, C.; Burrell, G.; Torriero, A.A.J.; Separovic, F.; Dunlop, N.F.; MacFarlane, D.R.; Bond, A.M. Electrochemistry of room temperature protic ionic liquids. *Journal of Physical Chemistry B* **2008**, *112*, 6923-6936.

24. Plechkova, N.V.; Seddon, K.R. Applications of ionic liquids in the chemical industry. *Chemical Society Reviews* **2008**, *37*, 123-150.
25. Xue, H.; Gao, Y.; Twamley, B.; Shreeve, J.M. New energetic salts based on nitrogen-containing heterocycles. *Chemistry of Materials* **2005**, *17*, 191-198.
26. MacFarlane, D.R.; Forsyth, M.; Izgorodina, E.I.; Abbott, A.P.; Annat, G.; Fraser, K. On the concept of ionicity in ionic liquids. *Physical Chemistry Chemical Physics* **2009**, *11*, 4962-4967.
27. Bradaric, C.J.; Downard, A.; Kennedy, C.; Robertson, A.J.; Zhou, Y.H. Industrial preparation of phosphonium ionic liquids. *Green Chem.* **2003**, *5*, 143-152.
28. Ermolaev, V.; Miluykov, V.; Rizvanov, I.; Krivolapov, D.; Zvereva, E.; Katsyuba, S.; Sinyashin, O.; Schmutzler, R. Phosphonium ionic liquids based on bulky phosphines: synthesis, structure and properties. *Dalton Transactions* **2010**, *39*, 5564-5571.
29. Busi, S.; Lahtinen, M.; Mansikkamaki, H.; Valkonen, J.; Rissanen, K. Synthesis, characterization and thermal properties of small R₂R' N⁺X⁻-type quaternary ammonium halides. *Journal of Solid State Chemistry* **2005**, *178*, 1722-1737.
30. MacFarlane, D.R.; Golding, J.; Forsyth, S.; Forsyth, M.; Deacon, G.B. Low viscosity ionic liquids based on organic salts of the dicyanamide anion. *Chemical Communications* **2001**, 1430-1431.
31. MacFarlane, D.R.; Pringle, J.M.; Johansson, K.M.; Forsyth, S.A.; Forsyth, M. Lewis base ionic liquids. *Chemical Communications* **2006**, 1905-1917.
32. Ercole, F.; Davis, T.P.; Evans, R.A. Photo-responsive systems and biomaterials: photochromic polymers, light-triggered self-assembly, surface modification, fluorescence modulation and beyond. *Polymer Chemistry* **2010**, *1*, 37-54.
33. Mortimer, R.J. Electrochromic Materials. In *Annual Review of Materials Research, Vol 41*, Clarke, D.R.F.P., Ed. 2011; Vol. 41, pp 241-268.
34. Zhang, H.X.; Meng, X.; Li, P. Light and thermal-stimuli responsive materials. *Progress in Chemistry* **2008**, *20*, 657-672.
35. Einaga, Y. Photo-switching magnetic materials. *Journal of Photochemistry and Photobiology C-Photochemistry Reviews* **2006**, *7*, 69-88.
36. Roy, D.; Cambre, J.N.; Sumerlin, B.S. Future perspectives and recent advances in stimuli-responsive materials. *Progress in Polymer Science* **2010**, *35*, 278-301.
37. Scarmagnani, S.; Walsh, Z.; Slater, C.; Alhashimy, N.; Paull, B.; Macka, M.; Diamond, D. Polystyrene bead-based system for optical sensing using spiropyran photoswitches. *Journal of Materials Chemistry* **2008**, *18*, 5063-5071.
38. Stitzel, S.; Byrne, R.; Diamond, D. LED switching of spiropyran-doped polymer films. *Journal of Materials Science* **2006**, *41*, 5841-5844.
39. Radu, A.; Byrne, R.; Alhashimy, N.; Fusaro, M.; Scarmagnani, S.; Diamond, D. Spiropyran-based reversible, light-modulated sensing with reduced photofatigue. *Journal of Photochemistry and Photobiology a-Chemistry* **2009**, *206*, 109-115.
40. Earle, M.J.; Gordon, C.M.; Plechkova, N.V.; Seddon, K.R.; Welton, T. *Anal. Chem.* **2007**, *79*, 758.
41. Miljanic, S.; Frkanec, L.; Meic, Z.; Zinic, M. Photoinduced Gelation by Stilbene Oxalyl Amide Compounds. *Langmuir* **2005**, *21*, 2754-2760.

42. Sanchez, A.M.; Barra, M.; de Rossi, R.H. On the mechanism of the acid/base-catalyzed thermal cis-trans isomerization of methyl orange. *Journal of Organic Chemistry* **1999**, *64*, 1604-1609.
43. Bazarnik, M.; Henzl, J.; Czajka, R.; Morgenstern, K. Light driven reactions of single physisorbed azobenzenes. *Chemical Communications* **2011**, *47*, 7764-7766.
44. Ichimura, K.; Oh, S.K.; Nakagawa, M. Light-driven motion of liquids on a photoresponsive surface. *Science* **2000**, *288*, 1624-1626.
45. Hanabusa, K.; Hiratsuka, K.; Kimura, M.; Shirai, H. Easy preparation and useful character of organogel electrolytes based on low molecular weight gelator. *Chemistry of Materials* **1999**, *11*, 649-655.
46. Terech, P.; Weiss, R.G. Low molecular mass gelators of organic liquids and the properties of their gels. *Chem. Rev.* **1997**, *97*, 3133-3159.
47. Murata, K.; Aoki, M.; Suzuki, T.; Harada, T.; Kawabata, H.; Komori, T.; Ohseto, F.; Ueda, K.; Shinkai, S. Thermal and Light Control of the Sol-Gel Phase Transition in Cholesterol-Based Organic Gels. Novel Helical Aggregation Modes As Detected by Circular Dichroism and Electron Microscopic Observation. *J. Am. Chem. Soc.* **1994**, *116*, 6664-6676.
48. Ahmed, S.A.; Sallenave, X.; Fages, F.; Mieden-Gundert, G.; Mueller, W.M.; Mueller, U.; Voegtle, F.; Pozzo, J.-L. Multiaddressable Self-Assembling Organogelators Based on 2H-Chromene and N-Acyl-1,w-amino Acid Units. *Langmuir* **2002**, *18*, 7096-7101.
49. Minkin, V.I. Photo-, thermo-, solvato-, and electrochromic spiroheterocyclic compounds. *Chem. Rev.* **2004**, *104*, 2751-2776.
50. Rosario, R.; Gust, D.; Garcia, A.A.; Hayes, M.; Taraci, J.L.; Clement, T.; Dailey, J.W.; Picraux, S.T. Lotus effect amplifies light-induced contact angle switching. *Journal of Physical Chemistry B* **2004**, *108*, 12640-12642.
51. Athanassiou, A.; Lygeraki, M.I.; Pisignano, D.; Lakiotaki, K.; Varda, M.; Mele, E.; Fotakis, C.; Cingolani, R.; Anastasiadis, S.H. Photocontrolled variations in the wetting capability of photochromic polymers enhanced by surface nanostructuring. *Langmuir* **2006**, *22*, 2329-2333.
52. Rosario, R.; Gust, D.; Hayes, M.; Jahnke, F.; Springer, J.; Garcia, A.A. Photon-modulated wettability changes on spiropyran-coated surfaces. *Langmuir* **2002**, *18*, 8062-8069.
53. Radu, A.; Scarmagnani, S.; Byrne, R.; Slater, C.; Lau, K.T.; Diamond, D. Photonic modulation of surface properties: a novel concept in chemical sensing. *Journal of Physics D-Applied Physics* **2007**, *40*, 7238-7244.
54. Coleman, S.; Byrne, R.; Minkovska, S.; Diamond, D. Thermal reversion of spirooxazine in ionic liquids containing the [NTf₂]⁽⁻⁾ anion. *Physical Chemistry Chemical Physics* **2009**, *11*, 5608-5614.
55. Byrne, R.; Coleman, S.; Fraser, K.J.; Raduta, A.; MacFarlane, D.R.; Diamond, D. Photochromism of nitrobenzospiropyran in phosphonium based ionic liquids. *Physical Chemistry Chemical Physics* **2009**, *11*, 7286-7291.
56. Wagner, K.; Byrne, R.; Zanoni, M.; Gambhir, S.; Dennany, L.; Breukers, R.; Higgins, M.; Wagner, P.; Diamond, D.; Wallace, G.G.; et al. A Multiswitchable Poly(terthiophene) Bearing a Spiropyran Functionality: Understanding Photo- and Electrochemical Control. *Journal of the American Chemical Society* **2011**, *133*, 5453-5462.

57. Szilagyi, A.; Sumaru, K.; Sugiura, S.; Takagi, T.; Shinbo, T.; Zrinyi, M.; Kanamori, T. Rewritable microrelief formation on photoresponsive hydrogel layers. *Chemistry of Materials* **2007**, *19*, 2730-2732.
58. Byrne, R.; Ventura, C.; Lopez, F.B.; Walther, A.; Heise, A.; Diamond, D. Characterisation and analytical potential of a photo-responsive polymeric material based on spiropyran. *Biosensors & Bioelectronics* **2010**, *26*, 1392-1398.
59. Brazel, C.S.; Peppas, N.A. SYNTHESIS AND CHARACTERIZATION OF THERMOMECHANICALLY AND CHEMOMECHANICALLY RESPONSIVE POLY(N-ISOPROPYLACRYLAMIDE-CO-METHACRYLIC ACID) HYDROGELS. *Macromolecules* **1995**, *28*, 8016-8020.
60. Schild, H.G.; Tirrell, D.A. INTERACTION OF POLY(N-ISOPROPYLACRYLAMIDE) WITH SODIUM NORMAL-ALKYL SULFATES IN AQUEOUS-SOLUTION. *Langmuir* **1991**, *7*, 665-671.
61. Bakker, E.; Buehlmann, P.; Pretsch, E. Carrier-Based Ion-Selective Electrodes and Bulk Optodes. 1. General Characteristics. *Chem. Rev.* **1997**, *97*, 3083-3132.
62. Bakker, E.; Xu, A.P.; Pretsch, E. Optimum Composition of Neutral Carrier Based Ph Electrodes. *Analytica Chimica Acta* **1994**, *295*, 253-262.
63. Malon, A.; Radu, A.; Qin, W.; Qin, Y.; Ceresa, A.; Maj-Zurawska, M.; Bakker, E.; Pretsch, E. Improving the detection limit of anion-selective electrodes: An iodide-selective membrane with a nanomolar detection limit. *Analytical Chemistry* **2003**, *75*, 3865-3871.
64. Sutter, J.; Radu, A.; Peper, S.; Bakker, E.; Pretsch, E. Solid-contact polymeric membrane electrodes with detection limits in the subnanomolar range. *Analytica Chimica Acta* **2004**, *523*, 53-59.
65. Ceresa, A.; Radu, A.; Peper, S.; Bakker, E.; Pretsch, E. Rational design of potentiometric trace level ion sensors. A Ag⁺-selective electrode with a 100 ppt detection limit. *Analytical Chemistry* **2002**, *74*, 4027-4036.
66. Del Sesto, R.E.; Corley, C.; Robertson, A.; Wilkes, J.S. Tetraalkylphosphonium-based ionic liquids. *Journal of Organometallic Chemistry* **2005**, *690*, 2536-2542.
67. Peng, B.; Zhu, J.W.; Liu, X.J.; Qin, Y. Potentiometric response of ion-selective membranes with ionic liquids as ion-exchanger and plasticizer. *Sensors and Actuators B-Chemical* **2008**, *133*, 308-314.
68. Shvedene, N.V.; Chernyshov, D.V.; Khrenova, M.G.; Formanovsky, A.A.; Baulin, V.E.; Pletnev, I.V. Ionic liquids plasticize and bring ion-sensing ability to polymer membranes of selective electrodes. *Electroanalysis* **2006**, *18*, 1416-1421.
69. Buehlmann, P.; Pretsch, E.; Bakker, E. Carrier-based ion-selective electrodes and bulk optodes. 2. Ionophores for potentiometric and optical sensors. *Chem. Rev.* **1998**, *98*, 1593-1687.
70. Chernyshov, D.V.; Shvedene, N.V.; Antipova, E.R.; Pletnev, I.V. Ionic liquid-based miniature electrochemical sensors for the voltammetric determination of catecholamines. *Anal. Chim. Acta* **2008**, *621*, 178-184.
71. Nadherná, M.; Opekar, F.; Reiter, J. Ionic liquid-polymer electrolyte for amperometric solid-state NO₂ sensor. *Electrochimica Acta* **2011**, *56*, 5650-5655.

72. Bird, C.L.; Kuhn, A.T. Electrochemistry of the Viologens. *Chemical Society Reviews* **1981**, *10*, 49-82.
73. Mortimer, R.J.; Dyer, A.L.; Reynolds, J.R. Electrochromic organic and polymeric materials for display applications. *Displays* **2006**, *27*, 2-18.
74. Rosseinsky, D.R.; Mortimer, R.J. Electrochromic systems and the prospects for devices. *Advanced Materials* **2001**, *13*, 783-793.
75. Wishart, J.F. Energy applications of ionic liquids. *Energy & Environmental Science* **2009**, *2*, 956-961.
76. Ahmad, S.; Deepa, M. Ionogels encompassing ionic liquid with liquid like performance preferable for fast solid state electrochromic devices. *Electrochem. Commun.* **2007**, *9*, 1635-1638.
77. Kavanagh, A.; Fraser, K.; Byrne, R.; Diamond, D. An electrochromic ionic liquid: design, characterisation and incorporation into a solid-state device. *Chemistry of Materials* **2011**, Submitted.
78. Deepa, M.; Ahmad, S.; Sood, K.N.; Alam, J.; Ahmad, S.; Srivastava, A.K. Electrochromic properties of polyaniline thin film nanostructures derived from solutions of ionic liquid/polyethylene glycol. *Electrochimica Acta* **2007**, *52*, 7453-7463.
79. Ahmad, S.; Singh, S. Electrochromic device based on carbon nanotubes functionalized poly(methyl pyrrole) synthesized in hydrophobic ionic liquid medium. *Electrochemistry Communications* **2008**, *10*, 895-898.
80. Deepa, M.; Awadhia, A.; Bhandari, S. Electrochemistry of poly(3,4-ethylenedioxythiophene)-polyaniline/Prussian blue electrochromic devices containing an ionic liquid based gel electrolyte film. *Physical Chemistry Chemical Physics* **2009**, *11*, 5674-5685.
81. Oehme, I.; Wolfbeis, O.S. Optical sensors for determination of heavy metal ions. *Mikrochimica Acta* **1997**, *126*, 177-192.
82. Zhu, J.W.; Zhai, J.Y.; Li, X.; Qin, Y. Applications of hydrophobic room temperature ionic liquids in ion-selective optodes. *Sensors and Actuators B-Chemical* **2011**, *159*, 256-260.
83. Topal, S.Z.; Ertekin, K.; Gurek, A.G.; Yenigul, B.; Ahsen, V. Tuning pH sensitivities of zinc phthalocyanines in ionic liquid modified matrices. *Sensors and Actuators B-Chemical* **2011**, *156*, 236-244.
84. Lunstroot, K.; Driesen, K.; Nockemann, P.; Gorller-Walrand, C.; Binnemans, K.; Bellayer, S.; Le Bideau, J.; Vioux, A. Luminescent ionogels based on europium-doped ionic liquids confined within silica-derived networks. *Chemistry of Materials* **2006**, *18*, 5711-5715.
85. Lunstroot, K.; Driesen, K.; Nockemann, P.; Van Hecke, K.; Van Meervelt, L.; Goerller-Walrand, C.; Binnemans, K.; Bellayer, S.; Viau, L.; Le Bideau, J.; et al. Lanthanide-doped luminescent ionogels. *Dalton Transactions* **2009**, 298-306.
86. Cheminet, N.; Jarrosson, T.; Lere-Porte, J.-P.; Serein-Spirau, F.; Cury, L.; Moreau, J.; Viau, L.; Vioux, A. One pot synthesis of fluorescent pi-conjugated materials: immobilization of phenylene-ethynylene polyelectrolytes in silica confined ionogels. *Journal of Materials Chemistry* **2011**, *21*, 13588-13593.
87. Wassercheid, P.; Welton, T. *Ionic liquids in synthesis*. Wiley-VCH: 2003.

88. Liu, Y.; Wang, M.; Li, J.; Li, Z.; He, P.; Liu, H.; Li, J. Highly active horseradish peroxidase immobilized in 1-butyl-3-methylimidazolium tetrafluoroborate room-temperature ionic liquid based sol-gel host materials. *Chem. Commun.* **2005**, 1778-1780.
89. Ohno, H.; Suzuki, C.; Fukumoto, K.; Yoshizawa, M.; Fujita, K. Electron Transfer Process of Poly(ethylene oxide)-Modified Cytochrome c in Imidazolium Type Ionic Liquid. *Chem. Lett.* **2003**, 32, 450-451.
90. Baker, S.N.; McCleskey, T.M.; Pandey, S.; Baker, G.A. Fluorescence studies of protein thermostability in ionic liquids. *Chem. Commun.* **2004**, 940-941.
91. Pernak, A.; Iwanik, K.; Majewski, P.; Grzymislawski, M.; Pernak, J. Ionic liquids as an alternative to formalin in histopathological diagnosis. *Acta Histochemica* **2005**, 107, 149-156.
92. Fujita, K.; MacFarlane, D.R.; Forsyth, M. Protein solubilising and stabilising ionic liquids. *Chem. Commun.* **2005**, 4804-4806.
93. Laszlo, J.A.; Compton, D.L. Comparison of peroxidase activities of hemin, cytochrome c and microperoxidase-11 in molecular solvents and imidazolium-based ionic liquids. *J. Mol. Catal. B: Enzym.* **2002**, 18, 109-120.
94. Yang, Z.; Pan, W. Ionic liquids: Green solvents for nonaqueous biocatalysis. *Enzyme Microb. Technol.* **2005**, 37, 19-28.
95. Zhao, H. Methods for stabilizing and activating enzymes in ionic liquids—a review. *J. Chem. Tech. Biotech* **2010**, 85, 891-907.
96. Abe, Y.; Yoshiyama, K.; Yagi, Y.; Hayase, S.; Kawatsura, M.; Itoh, T. A rational design of phosphonium salt type ionic liquids for ionic liquid coated-lipase catalyzed reaction. *Green Chemistry* **2010**, 12, 1976-1980.
97. Zhang, Y.; Zheng, J. Direct electrochemistry and electrocatalysis of myoglobin immobilized in hyaluronic acid and room temperature ionic liquids composite film. *Electrochem. Commun.* **2008**, 10, 1400-1403.
98. Zhang, J.; Lei, J.; Liu, Y.; Zhao, J.; Ju, H. Highly sensitive amperometric biosensors for phenols based on polyaniline-ionic liquid-carbon nanofiber composite. *Biosens. Bioelectron.* **2009**, 24, 1858-1863.
99. Torimoto, T.; Tsuda, T.; Okazaki, K.-i.; Kuwabata, S. New Frontiers in Materials Science Opened by Ionic Liquids. *Adv. Mater.* **2010**, 22, 1196-1221.
100. Ohno, K.-i.; Tachikawa, K.; Manz, A. Microfluidics: Applications for analytical purposes in chemistry and biochemistry. *ELECTROPHORESIS* **2008**, 29, 4443-4453.
101. Brady, S.; Dunne, L.E.; Lynch, A.; Smyth, B.; Diamond, D. *Personalised Health Management Systems: The Integration of Innovative Sensing, Textile, Information and Communication Technologies*. IOS Press: 2005; Vol. 117/2005.
102. Bhandari, P.; Narahari, T.; Dendukuri, D. 'Fab-Chips': a versatile, fabric-based platform for low-cost, rapid and multiplexed diagnostics. *Lab on a Chip* **2011**, 11, 2493-2499.
103. Schmid-Wendtner, M.H.; Korting, H.C. The pH of the skin surface and its impact on the barrier function. *Skin Pharmacol Physiol* **2006**, 19, 296-302.
104. Patterson, M.J.; Galloway, S.D.R.; Nimmo, M.A. Effect of induced metabolic alkalosis on sweat composition in men. *Acta Physiol. Scand.* **2002**, 174, 41-46.

105. Granger, D.; Marsolais, M.; Burry, J.; Laprade, R. Na⁺/H⁺ exchangers in the human eccrine sweat duct. *American Journal of Physiology-Cell Physiology* **2003**, 285, C1047-C1058.
106. Patterson, M.J.; Galloway, S.D.R.; Nimmo, M.A. Variations in regional sweat composition in normal human males. *Exp Physiol* **2000**, 85, 869-875.
107. Morgan, R.M.; Patterson, M.J.; Nimmo, M.A. Acute effects of dehydration on sweat composition in men during prolonged exercise in the heat. *Acta Physiol. Scand.* **2004**, 182, 37-43.
108. O'Toole, M.; Shepherd, R.; Wallace, G.G.; Diamond, D. Inkjet printed LED based pH chemical sensor for gas sensing. *Anal. Chim. Acta* **2009**, 652, 308-314.
109. Smyth, C.n.; Lau, K.T.; Shepherd, R.L.; Diamond, D.; Wu, Y.; Spinks, G.M.; Wallace, G.G. Self-maintained colorimetric acid/base sensor using polypyrrole actuator. *Sensors and Actuators B: Chemical* **2008**, 129, 518-524.
110. Benito-Lopez, F.; Coyle, S.; Byrne, R.; O'Toole, C.; Barry, C.; Diamond, D. Simple Barcode System Based on Inonogels for Real Time pH-Sweat Monitoring. *Body Sensor Networks (BSN), 2010 International Conference on* **2010**, 291.
111. Coyle, S.; Benito-Lopez, F.; Radu, T.; Lau, K.T.; Diamond, D. Fibers and fabrics for chemical and biological sensing. *J. Text. App.* **2010**, 14, 64.
112. Nilsson, D.; Kugler, T.; Svensson, P.-O.; Berggren, M. An all-organic sensor-transistor based on a novel electrochemical transducer concept printed electrochemical sensors on paper. *Sensors and Actuators B: Chemical* **2002**, 86, 193-197.
113. Chaubey, A.; Pande, K.K.; Singh, V.S.; Malhotra, B.D. Co-immobilization of lactate oxidase and lactate dehydrogenase on conducting polyaniline films. *Anal. Chim. Acta* **2000**, 407, 97-103.
114. Weber, J.; Kumar, A.; Kumar, A.; Bhansali, S. Novel lactate and pH biosensor for skin and sweat analysis based on single walled carbon nanotubes. *Sensors and Actuators B: Chemical* **2006**, 117, 308-313.
115. Yashina, E.I.; Borisova, A.V.; Karyakina, E.E.; Shchegolikhina, O.I.; Vagin, M.Y.; Sakharov, D.A.; Tonevitsky, A.G.; Karyakin, A.A. Sol,àíGel Immobilization of Lactate Oxidase from Organic Solvent: Toward the Advanced Lactate Biosensor. *Anal. Chem.* **2010**, 82, 1601-1604.
116. Cai, X.; Yan, J.; Chu, H.; Wu, M.; Tu, Y. An exercise degree monitoring biosensor based on electrochemiluminescent detection of lactate in sweat. *Sensors and Actuators B: Chemical* **2010**, 143, 655-659.
117. Wang, Y.; Xu, H.; Zhang, J.M.; Li, G. Electrochemical Sensors for Clinic Analysis. *Sensors* **2008**, 8, 2043-2081.
118. Weil, M.H.; Afifi, A.A. Experimental and Clinical Studies on Lactate and Pyruvate as Indicators of the Severity of Acute Circulatory Failure (Shock). *Circulation* **1970**, 41, 989-1001.
119. Green, J.M.; Pritchett, R.C.; Crews, T.R.; McLester, J.R.; Tucker, D.C. Sweat lactate response between males with high and low aerobic fitness. *European Journal of Applied Physiology* **2004**, 91, 1-6.
120. Billat, L.V. Use of blood lactate measurements for prediction of exercise performance and for control of training: Recommendations for long-distance running. *Sports Med.* **1996**, 22, 157-175.
121. Branco, L.C.; Pina, F. Intrinsically photochromic ionic liquids. *Chemical Communications* **2009**, 6204-6206.

122. Zhang, Q.; Zhang, S.; Liu, S.; Ma, X.; Lu, L.; Deng, Y. Ionic liquid-modified dyes and their sensing performance toward acids in aqueous and non-aqueous solutions. *Analyst* **2011**, *136*, 1302-1304.
123. Branco, A.; Branco, L.C.; Pina, F. Electrochromic and magnetic ionic liquids. *Chem. Commun.* **2011**, *47*, 2300-2302.