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1 **Optical absorption and EPR studies on tenorite mineral**

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13  
14 **Key words:** Tenorite mineral, EPR spectrum, Optical absorption spectrum, Octahedral  
15 symmetry, Cu(II),  $d^9$  structure,  $d^2$  structure.

16  
17 **Abstract**

18  
19 Optical absorption and EPR studies of the mineral tenorite, a cupric oxide, which  
20 originated from Mexico and contains 54.40 wt% of CuO. EPR spectral results indicate  
21 two Cu(II) closely interacting ions to give a  $d^2$  type structure. The calculated spin  
22 Hamiltonian at Rt and LNT are  $g = 2.160$  and  $D = 125$  G. The intensity of resonance line  
23 is not the same in low and high field regions. The optical absorption spectrum is due to  
24 Cu(II) which three sets of energies indicating Cu(II) in two independent tetragonal  $C_{4v}$   
25 symmetry, in addition to  $d^2$  structure of octahedral coordination. The octahedral and  
26 tetragonal field parameters are compared with those reported for several other copper  
27 containing minerals.

28  
29  
30 **Introduction**

31

32           Among the most important metals, copper plays a vital role in various fields of  
33 science. Naturally this metal appears in various forms. Among them cupric oxide  
34 (CuO) is one. In nature, it occurs in the form of a mineral called tenorite. It is  
35 monoclinic with  $C_{2h}^6(C2/C)$  point group. There are four CuO molecules in the unit  
36 cell and two CuO units in the primitive cell and thus the copper and oxygen occupy  $C_1$   
37 and  $C_2$  symmetry positions respectively. Each Cu atom is situated at the center of the  
38 oxygen parallelogram. Each oxygen atom has a distorted tetrahedral copper atom [1-2].  
39 Since CuO has been studied extensively towards super conducting material, a few  
40 references have been mentioned [3-5]. In the first paper [3], a series of monomer, dimer  
41 and tetramer copper complexes have been prepared. However, the results indicate a  
42 broad line and nothing about EPR zero field splitting parameter. It is also mentioned that  
43 the copper aggregates giving the EPR resonance are magnetically isolated from the main  
44 CuO cluster. The second paper [4] discusses the role of charge stripes in CuO to  
45 understand the super conducting nature. The third [5] paper deals the multi sub lattice  
46 structure of CuO. Tenorite is an interesting mineral. Moreover no electron  
47 paramagnetic resonance (EPR) and optical absorption studies so far were reported for the  
48 tenorite mineral. Divalent copper being a  $d^9$  system is prone to Jahn-Teller distortion and  
49 EPR study of this ion will give valuable information about the type of distortion. The  
50 optical absorption spectral studies of transition metal complexes provide information on  
51 excited state energies, symmetrical and structural distortions. Therefore the authors  
52 propose to report the results of spectral studies by EPR and optical absorption techniques  
53 of tenorite mineral.

54

## 55 **Experimental**

56

57           Tenorite mineral originated from Mexico donated by Dr.Michel Delienes, is used  
58 in the present work. It is dark brownish black in colour. The copper content is  
59 determined by standard spectrophotometric method [6]. EPR spectra of the sample in  
60 powder form are recorded both at room temperature (RT) and liquid nitrogen temperature  
61 (LNT) on a Varian E-112 EPR spectrometer operating at X - band frequencies ( $\nu = 9.07$   
62 GHz) having 100 kHz field modulation and phase sensitive detection to obtain first

63 derivative spectra. The optical absorption spectra of the sample are recorded at room  
64 temperature on Carey 5E UV-Vis-NIR spectrophotometer in mull form in the range 200 -  
65 2500 nm.

66

## 67 **Theory**

68

69 Cu(II) has the electronic configuration  $[\text{Ar}]3d^9$ , where  $[\text{Ar}]$  stands for the closed  
70 argon shell. To explain the observed EPR and optical absorption spectra the following  
71 options are kept in mind. (i) If the crystal contains two types of independent Cu(II) ions,  
72 this will lead to the arrangement of  $d^9$  configuration. (ii) If two Cu(II) ions interact to  
73 form two unpaired electrons, this arrangement leads to  $d^2$  or  $d^0$  configuration. For  $d^2$   
74 configuration, ground state term is  $^3F$  and excited state terms  $^3P, ^1G, ^1D$  and  $^1S$ . The  
75 transitions from the ground state to the three states  $^1G, ^1D$  and  $^1S$  are spin forbidden and  
76 are very weak. The two remaining  $^3F$  and  $^3P$  can have spin-permitted transitions. In  
77 octahedral field,  $^3P$  transformed to  $^3T_{1g}(\text{P})$  state and  $^3F$  split into  $^3A_{2g}(\text{F}), ^3T_{1g}(\text{F})$  and  
78  $^3T_{2g}(\text{F})$  states. Of these  $^3T_{1g}(\text{F})$  state lies lowest according to Hund's rule and becomes  
79 ground state [7].  $d^0$  being diamagnetic, no EPR or optical spectra are expected. On the  
80 other hand, in an octahedral crystal field,  $d^9$  ion has the corresponding ground state  
81 electronic configuration as  $t_{2g}^6 e_g^3$  and yields  $^2E_g$  term. The excited electronic  
82 configuration  $t_{2g}^5 e_g^4$  corresponds to  $^2T_{2g}$  term. Thus only one single electronic transition  
83  $^2E_g$  to  $^2T_{2g}$  is expected in an octahedral crystal field. Normally the ground  $^2E_g$  state splits  
84 due to Jahn-Teller effect and hence lowering of symmetry is expected for Cu(II) ion.  
85 This state splits into  $^2B_{1g}(d_{x^2-y^2})$  and  $^2A_{1g}(d_{x^2})$  states in tetragonal symmetry and the  
86 excited term  $^2T_{2g}$  also splits into  $^2B_{2g}(d_{xy})$  and  $^2E_g(d_{xz}, d_{yz})$  levels. In rhombic field  $^2E_g$   
87 ground state splits into  $^2A_{1g}(d_{x^2-y^2})$  and  $^2A_{2g}(d_{z^2})$  where as  $^2T_{2g}$  splits into  $^2B_{1g}(d_{xy})$   
88 and  $^2B_{2g}(d_{xz})$  and  $^2B_{3g}(d_{yz})$  states. Thus three bands are expected for tetragonal ( $C_{4v}$ )  
89 symmetry and four bands are expected for rhombic ( $D_{2h}$ ) symmetry [8].

90

91

92

93 **Results and analysis**

94

95 **EPR Studies**

96

97 A poly crystalline sample of tenorite mineral is taken in a quartz EPR tube for  
98 room temperature measurements. EPR spectrum obtained at room temperature is given  
99 in Fig. 1 marked with RT. A look at the spectrum immediately points that we are dealing  
100 with a system, whose spin value is not equal to  $\frac{1}{2}$ . Generally, a polycrystalline sample of  
101 copper ion gives a single broad line due to dipolar-dipolar broadening, around  $g \sim 2.3$ .  
102 However, the present spectrum is completely different.

103

104 Cu(II) being  $d^9$  has one unpaired electron. If another Cu(II) is present near by  
105 having proper Cu-O-Cu angle with Cu-Cu bond, there exists a possibility of interaction  
106 between these two ions. If the two ions interact, it will result  $S = 0$  (diamagnetic, paired,  
107 singlet) or  $S = 1$  (paramagnetic, unpaired, triplet). As a well-resolved EPR spectrum is  
108 observed at room temperature, we can suggest unambiguously that we are having a triplet  
109 state EPR spectrum. For a system with  $S = 1$ , the electronic levels splits into  $|\pm 1\rangle$  and  $|0\rangle$   
110 in the absence of an applied magnetic field with separation equal to  $D$  ( $D$  being zero field  
111 splitting constant). When an external magnetic field is applied, these two levels split into  
112 three levels,  $|+1\rangle$ ,  $|0\rangle$  and  $|-1\rangle$  and one can observe two transitions, separated by  $2D$  [9].

113

114 The room temperature EPR spectrum of tenorite mineral can be explained by  
115 considering interaction between two copper ions in the structure. The structure contains  
116 two types of O-Cu-O angles ( $84.5$  and  $95.5^\circ$ ): copper has four oxygen neighbours at  $1.96$   
117 Å, the next nearest neighbours being two oxygens at  $2.78$  Å. The shortest Cu-Cu  
118 distance is  $2.90$  Å [2]. This close distance between the two copper ions is favorable for  
119 interaction between the ions. The calculated values are  $g = 2.160$ ,  $D = 125$  G. The  
120 resonance line on low field side is slightly more intense than the one on high field side.  
121 However, the half field line at  $g \sim 4.0$  is not observed. This may be due to its low  
122 intensity (forbidden transition) or weak interaction with very low exchange coupling

123 constant (J). The information about J needs extensive magnetic moment measurements at  
124 various temperatures down to 77 K [10].

125

126 The sample is cooled to 77 K and the EPR spectrum thus obtained is given in Fig. 1  
127 marked with LNT. This is also very similar to RT spectrum. The calculated g and D  
128 values are 2.160 and 125 G. These values match fairly well with RT values. Here also,  
129 half field transitions could not be seen. However, the resonance line at low field is less  
130 intense than the one on high field side. This indicates that the ground state is  $|-1\rangle$ .  
131 However, we did not notice any paramagnetic- antiferromagnetic phase transition, as  
132 observed in a few copper compounds [1]

133

### 134 **Optical absorption studies**

135

136 The optical absorption spectrum of tenorite mineral recorded in mull form at room  
137 temperature (RT) is shown in Fig. 2. The spectrum observed here has bands at  
138 41670,35090,23250,18690,17850,16670,15530,12930,12170, 10310 and 8865  $\text{cm}^{-1}$ . For  
139 easy analysis of the spectrum the bands are divided into three sets (a) 8865,10310, 12930  
140  $\text{cm}^{-1}$ ,(b) 12170,16670, 18690  $\text{cm}^{-1}$  and (c) 15530, 17850, 23250, 35090 and 41670  $\text{cm}^{-1}$ .  
141 Cu(II) being  $d^9$  has one unpaired electron and generally gives optical absorption energies  
142 around 8000,10000 and 14000  $\text{cm}^{-1}$  [11] and energies around 13000, 16000 and 18000  
143  $\text{cm}^{-1}$  [12] for octahedral geometry assuming  $C_{4v}$  site symmetry for Cu(II) ion in the  
144 sample. The bands observed at 8865, 10310 and 12930  $\text{cm}^{-1}$  are assigned to  ${}^2B_1 \rightarrow {}^2A_1$ ,  
145  ${}^2B_1 \rightarrow {}^2B_1$  and  ${}^2B_1 \rightarrow {}^2E$  transitions respectively. The octahedral (Dq) and tetragonal (Ds,  
146 Dt) field parameters are evaluated from the following expressions:

147

$$4Ds + 5Dt : {}^2B_1 \rightarrow {}^2A_1$$

148

$$10Dq : {}^2B_1 \rightarrow {}^2B_1$$

149

$$10Dq + 3Ds - 5Dt : {}^2B_1 \rightarrow {}^2E$$

150

151 The parameters evaluated are found to be  $Dq = 1031 \text{ cm}^{-1}$ ,  $Ds = 1640 \text{ cm}^{-1}$  and  $Dt = 461$   
152  $\text{cm}^{-1}$ . The another set of bands at 12170 16670 18690  $\text{cm}^{-1}$  are assigned to  ${}^2B_1 \rightarrow {}^2A_1$ ,  
153  ${}^2B_1 \rightarrow {}^2B_1$  and  ${}^2B_1 \rightarrow {}^2E$  transitions respectively. Using the above expressions the

154 parameters evaluated are found to be  $Dq = 1667 \text{ cm}^{-1}$ ,  $Ds = 2027 \text{ cm}^{-1}$  and  $Dt = 812 \text{ cm}^{-1}$ .  
155 The band positions along with their assignments are given in Table – 1. The values of  $Dq$   
156 and  $Dt$  are of the same sign in the present sample thus indicating tetragonally elongated  
157 octahedral site for Cu(II) ion [13]. The values of tetragonal parameters are higher in  
158 second set when compared with the first set. It indicates tetragonal distortion is more in  
159 second set than in first set. For comparison the octahedral and tetragonal field parameters  
160 with those reported earlier in different copper bearing minerals are presented in Table –2.  
161 The observed parameters in the present sample match fairly with other samples.  
162

163 As mentioned in EPR spectral analysis where we assumed a two unpaired  
164 electrons to explain the observed one, here also we assumed two electrons present in the  
165 sample. The optical absorption spectrum also has two strong bands at 15530 and 23250  
166  $\text{cm}^{-1}$ , a weak band at 17850  $\text{cm}^{-1}$  along with charge transfer bands at 35090 and 41670  
167  $\text{cm}^{-1}$ . Some previously reported spectra for  $d^2$  ion surrounded six oxygen ligands showed  
168 bands at 17400, 25200  $\text{cm}^{-1}$  [14] and 16900, 23500  $\text{cm}^{-1}$  [15]. The two stronger bands  
169 observed at 15530 and 23250  $\text{cm}^{-1}$  are assigned to the transitions  ${}^3T_{1g}(F) \rightarrow {}^3T_{2g}(F)$  and  
170  ${}^3T_{1g}(F) \rightarrow {}^3T_{1g}(P)$  respectively in an octahedral coordination. Using T-S diagram [16] for  
171  $d^2$  ion,  $10Dq$  is calculated to be 16800  $\text{cm}^{-1}$  and  $B = 600 \text{ cm}^{-1}$ . A third band could be  
172 expected to be charge transfer band [17]. Accordingly the band observed at 35090  $\text{cm}^{-1}$   
173 is attributed to  ${}^3T_{1g}(F) \rightarrow {}^3A_{2g}(F)$ . The lower singlet energy transition  ${}^3T_{1g}(F) \rightarrow {}^1E_g(D)$ ,  
174  ${}^1T_{2g}(D)$  is observed in the range 10200 – 12000  $\text{cm}^{-1}$ [17] and this transition is spin-  
175 forbidden and is very weak. But in the observed spectrum the bands at 12 170 and 10310  
176  $\text{cm}^{-1}$  are intense and are not attributed  $d^2$  ion. But the band observed at 17850  $\text{cm}^{-1}$  is  
177 attributed to  ${}^3T_{1g}(F) \rightarrow {}^1A_{1g}(G)$  transition of  $d^2$  ion. The unassigned band at 41670  $\text{cm}^{-1}$   
178 may be due to charge transfer transition between metal and ligand.

179

180

181

182 **Conclusions**

183

184 Tenorite is copper mineral. Its analysis indicates that CuO is 54.40 wt%. The  
185 optical absorption spectrum is due to Cu(II) in two independent positions in tetragonal  
186  $C_{4v}$  symmetry. Further it might be noticed that Cu(II) + Cu(II) are in a weaker interaction  
187 to form a  $d^2$  configuration. The energies are attributed to octahedral coordination of  
188 Cu(II) + Cu(II) with  $d^2$  configuration. EPR spectrum reveals the presence of only  $d^2$  ion  
189 but not  $d^9$  ion. The calculated values are  $g = 2.160$  and  $D = 125$  G. Even at low  
190 temperature no changes are observed in the EPR spectrum ruling out any phase transition.  
191 The intensity of resonance line is not uniform in low and high field regions and shows  
192 that the ground state is  $|-1\rangle$ . Thus the results of the optical and EPR investigations  
193 reported in the present communication conclusively prove that the site symmetry of  
194 Cu(II) + Cu(II) ( $d^2$ ) ion in tenorite mineral is octahedral coordination with  ${}^3T_{1g}(F)$  as  
195 ground state.

196

#### 197 **Acknowledgements**

198

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200 Dept. of Mineralogy, Institute Royal des Sciences Naturalles, de Belgiqui, Bruxelles for  
201 providing tenorite mineral sample.

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**Table – 1****Assignments of Cu(II) bands in tenorite mineral**

Transitions	Band positions	
	Observed Wave number	Calculated Wave number
Site 1		
${}^2B_1 \rightarrow {}^2A_1$	8865	8865
${}^2B_1 \rightarrow {}^2B_1$	10310	10310
${}^2B_1 \rightarrow {}^2E$	12930	12925
Site 2		
${}^2B_1 \rightarrow {}^2A_1$	12170	12168
${}^2B_1 \rightarrow {}^2B_1$	16670	16670
${}^2B_1 \rightarrow {}^2E$	18690	18691
Site 3		
${}^3T_{1g}(F) \rightarrow {}^3T_{2g}(F)$	15530	
${}^3T_{1g}(F) \rightarrow {}^3T_{1g}(P)$	23250	
${}^3T_{1g}(F) \rightarrow {}^3A_{2g}(F)$	35090	
${}^3T_{1g}(F) \rightarrow {}^1A_{1g}(G)$	17850	
Charge transfer	41670	

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**Table - 2**  
**Comparison of octahedral and tetragonal field parameters for Cu(II)**  
**in different Minerals**

Sl.No.	Name of the mineral	Dq (cm <sup>-1</sup> )	Ds (cm <sup>-1</sup> )	Dt(cm <sup>-1</sup> )	Reference
1	Melanterite	1053	1510	414	[18]
2	Pseudo-malachite	1149	1617	491	[19]
3	Connellite	1205	1525	450	[20]
4	Cornetite	1257	1434	380	[21]
5	Tenorite	1031	1640	461	Present work
6	Chrysocolla	1400	3000	800	[12]
7	Chalcanthite	1250	3240	710	[22]
8	Malachite	1250	3245	700	[22]
9	Tenorite	1667	2027	812	Present work

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226 **References**

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- 228 1. A.B.Kuzmenko, D.van der Marel, P.T.M. Van Bentum, E.A. Tishchenko, C.Presura  
229 and A.Bush, Physical Review B, **63**, 1 (2001).
- 230 2. A.F.Wells, "Structural Inorganic Chemistry", 5<sup>th</sup> Edition,  
231 Clarendon Press, Oxford (1987).
- 232 3. R. J. Singh, Alex Punnoose, Jilson Mathew, B.P. Mayura, M.Umar and M Haque,  
233 Phys. Reve. B **49**, 1346 (1994).
- 234 4. X.G. Zheng, C.N.Xu, Y.Tomokiyo, E.Tanaka, H.Yamada and Y.Soejima,  
235 Phys.Rev. Lett. **85**, 5170 (2000)
- 236 5. Koichi Kindo, Makoto Honda, Teruo Kohashi and Muneyuki Date,  
237 J. Phys. Soc. (Japan) **59**, 2332 (1990).
- 238 6. J.Basett, R.C.Denry, G.H.Geffery, J.Mendham: "Vogel's text book of Quantitative  
239 Inorganic Analysis" ELBS Edn. (1979).
- 240 7. S.F.A. Kettle "Coordination Compounds", ELBS Edition (1971).
- 241 8. M.A. Hitachmann, T.D. Waite, Inorg. Chem. **15**, 2150 (1976).
- 242 9. A. Abragam and B.Bleane, "EPR Transition Ions" Clarendon Press, Oxford (1970).
- 243 10. I.Sougandi, R.Mahalakshmi, R. Kannappan, T.M. Rajendrian, R.Venkatesan and  
244 P.Sambasiva Rao, Trans. Metal. Chem. **27**,512(2002).
- 245 11. P.Sreeramulu, K.M.Reddy, A.Sundar Jacob and B.J.Reddy, J.Crystallo. and Spectro.  
246 Res., **20**,1(1990).
- 247 12. S.V.J.Lakshman, B.J.Reddy, K.Janardanam, V.L.R.Moorthy and P. Rama Devi, .  
248 J.Phys. **B28**, 1122(1978).
- 249 13. B.J. Hathaway and D.E. Billing, Coord. Chem. Rev. **5**,143(1970).
- 250 14. W. Low, J. Physik, Chem. Frank., **13**, 107 (1957).
- 251 15. R.C. Furman and C.S. Garner., J. Am. Chem. Soc. **72**,1785 (1950).
- 252 16. Y.Tanabe and S. Sugano, J.Phys.Soc.Japan, **9**,753 (1954).
- 253 17. C.J. Ballhausen and F. Winther, Acta. Chem. Scand., **12**,1729(1959).
- 254 18. B.J.Reddy, S.N. Reddy, R.V.S.S.N. Ravikumar and P.S. Rao, ICTM-10 Conference,  
255 POII, Stuttgart, Germany (1995).
- 256 19. A.V. Chandrasekhar, M.Venkataramanaih, R.V.S.S.N. Ravikumar, B.J.Reddy and

- 257 Y.P.Reddy, Ind. J.Pure & Appl.Phys. **35**,71 (1997).
- 258 20. P.Sreeramulu, K.M.Reddy, A. Sunder Jacob and B.J.Reddy, J.Cryst and Spect.  
259 Research., **20**,93 (1990).
- 260 21. M.Venkataramanaiah, R.V.S.S.N. Ravikumar, G.Srinivasulu, B.J.Reddy and  
261 P.S. Rao Ferroelectrics, **175**,175 (1996).
- 262 22. S.V.J.Lakshman and B.J.Reddy, Can.Min. **12**,207 (1973).
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273 **List of Figures**

274

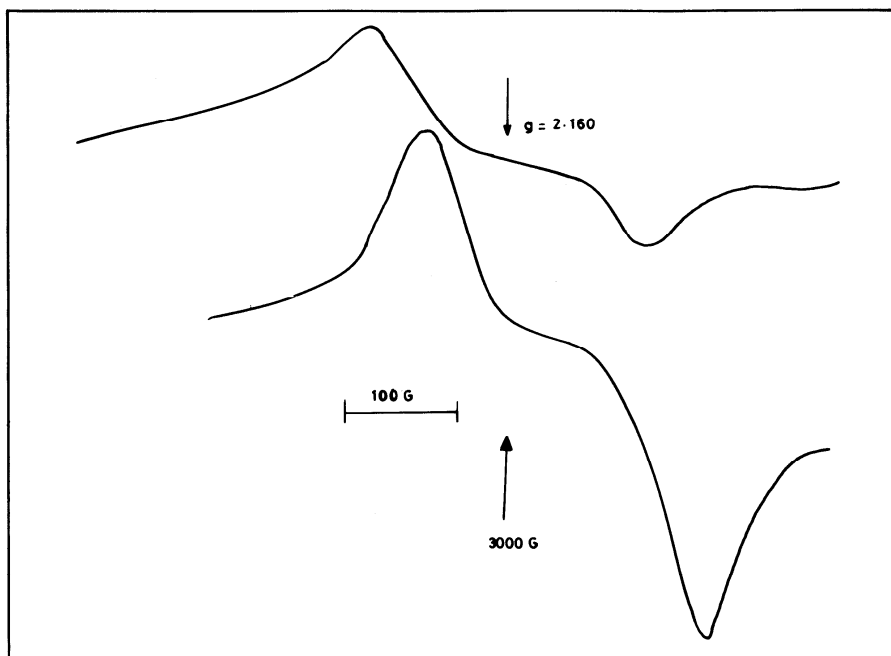
275 Fig. 1 EPR spectrum of tenorite mineral at room and liquid nitrogen temperature.

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277 Fig.2.Optical absorption spectrum of tenorite mineral at room temperature.

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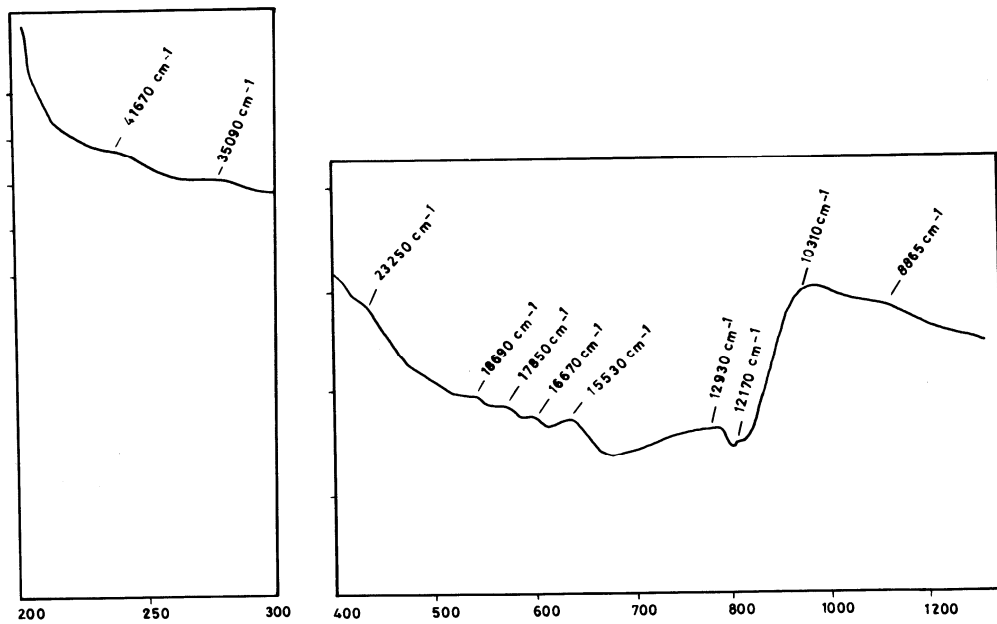
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282 **Figure 1**

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287 **Figure 2**

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